

What are Acids, Bases, and Salts?

The Swedish chemist Svante Arrhenius introduced the theory of ionization and used this theory to explain much about the behavior of acids and bases.

- An Arrhenius acid is defined as any compound that dissociates in aqueous solution to form H^+ ions.



- An Arrhenius base is defined as any compound that dissociates in aqueous solution to form OH^- ions.



- Salts are compounds that dissociate in aqueous solution releasing neither H^+ ions nor OH^- ions.



Using the **Arrhenius definition**, classify the following examples as acids, bases, or salts:

HBr	<u>acid</u>	KCl	<u>salt</u>
$Mg(OH)_2$	<u>base</u>	H_3PO_4	<u>acid</u>
HCl	<u>acid</u>	HClO	<u>acid</u>
KNO_3	<u>salt</u>	$Al(OH)_3$	<u>base</u>
HFO_4	<u>acid</u>	$KC_2H_3O_2$	<u>salt</u>
$Ba(OH)_2$	<u>base</u>	NaCl	<u>salt</u>

Note Taking Worksheet Acids Bases And Salts

L Reisser



Note Taking Worksheet Acids Bases And Salts:

Note Taking Worksheet Acids Bases And Salts Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Note Taking Worksheet Acids Bases And Salts**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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