

Nucleons:

- are particles occupying the nucleus
- consist of + charged protons and neutral charge neutrons
- have almost 2000 times the mass of electrons
- are made up of quark and leptons

Nuclear Notation:



$^{27}_{13}\text{Al}$ has 13 protons and 14 neutrons, for a total of 27 nucleons.

$^{28}_{13}\text{Al}$ has 13 protons and 15 neutrons, for a total of 28 nucleons.

isotopes: atoms of same element with different numbers of neutrons (different masses)

Forces Acting on Nucleons:

strong forces -

- forces of attraction between nucleons
- are independent of the charges of the nucleon
- are short range (exist only between closest neighbors)

electrical force -

- force of repulsion between + charged protons
- are long range

When are nuclei unstable?

- large nuclei ($Z > 82$) - electrical forces of repulsion are greater than strong forces of attraction
- wrong neutron : proton ratio

A radioactive isotope:

- has an unstable nucleus
- spontaneously emits a particle and charge into another element

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Table of Contents Note Taking Guide Episode 4answers Chemistry

1. Understanding the eBook Note Taking Guide Episode 4answers Chemistry
 - The Rise of Digital Reading Note Taking Guide Episode 4answers Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Note Taking Guide Episode 4answers Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Note Taking Guide Episode 4answers Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Note Taking Guide Episode 4answers Chemistry
 - Personalized Recommendations
 - Note Taking Guide Episode 4answers Chemistry User Reviews and Ratings
 - Note Taking Guide Episode 4answers Chemistry and Bestseller Lists
5. Accessing Note Taking Guide Episode 4answers Chemistry Free and Paid eBooks
 - Note Taking Guide Episode 4answers Chemistry Public Domain eBooks

- Note Taking Guide Episode 4answers Chemistry eBook Subscription Services
- Note Taking Guide Episode 4answers Chemistry Budget-Friendly Options
- 6. Navigating Note Taking Guide Episode 4answers Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Note Taking Guide Episode 4answers Chemistry Compatibility with Devices
 - Note Taking Guide Episode 4answers Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Note Taking Guide Episode 4answers Chemistry
 - Highlighting and Note-Taking Note Taking Guide Episode 4answers Chemistry
 - Interactive Elements Note Taking Guide Episode 4answers Chemistry
- 8. Staying Engaged with Note Taking Guide Episode 4answers Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Note Taking Guide Episode 4answers Chemistry
- 9. Balancing eBooks and Physical Books Note Taking Guide Episode 4answers Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Note Taking Guide Episode 4answers Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Note Taking Guide Episode 4answers Chemistry
 - Setting Reading Goals Note Taking Guide Episode 4answers Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Note Taking Guide Episode 4answers Chemistry
 - Fact-Checking eBook Content of Note Taking Guide Episode 4answers Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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