

Reaction Rates Test Name:

20 K 9 T 9 C 8 A

Assessment of Learning (Product)

Multiple Choice; Circle the most correct answer. (20K)

1. The kinetics of the reaction below were studied and it was determined that the reaction rate increased by a factor of 9 when the concentration of B was tripled. The reaction is _____ order in B.



- A) zero B) first C) second D) third E) one-half

2. Consider the following reaction:



If the rate of formation of NH_3 is 4.0×10^{-4} mol/s, then the rate of consumption of H_2 is:

- A) 2.0×10^{-4} mol/s B) 4.0×10^{-4} mol/s C) 6.0×10^{-4} mol/s D) 1.2×10^{-3} mol/s

3. Which of the following could be used to describe the rate of a reaction?

- A) $\frac{\text{change in time}}{\text{change in concentration}}$ B) $\frac{\text{change in mass}}{\text{change in concentration}}$
C) $\frac{\text{change in concentration}}{\text{change in time}}$ D) $\frac{\text{change in concentration}}{\text{change in mass}}$

4. Which of the following would **not** be units for reaction rate?

- A) g/s B) M/min C) kJ/mol D) mol/min

5. Consider the following reaction:



The activation energy of the forward reaction is 134 kJ. What is the activation energy for the reverse reaction?

- A) -134 kJ B) -100 kJ C) 234 kJ D) 368 kJ

6. Which of the following changes will increase the average kinetic energy of reactant molecules?

- A) adding a catalyst B) increasing the temperature
C) increasing the surface area D) increasing the concentration

7. The minimum amount of energy needed to start a reaction is called the:

- A) activation energy B) energy of reaction
C) entropy of reaction D) reaction mechanism energy

8. In general, a chemical reaction requiring a large activation energy will proceed:

- A) at a fast rate B) at a slow rate
C) only at low temperatures D) only at low concentrations

Reaction Rates Assessment Answers

O. Ray Whittington

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Table of Contents Reaction Rates Assessment Answers

1. Understanding the eBook Reaction Rates Assessment Answers
 - The Rise of Digital Reading Reaction Rates Assessment Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Rates Assessment Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reaction Rates Assessment Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Rates Assessment Answers
 - Personalized Recommendations
 - Reaction Rates Assessment Answers User Reviews and Ratings
 - Reaction Rates Assessment Answers and Bestseller Lists

5. Accessing Reaction Rates Assessment Answers Free and Paid eBooks
 - Reaction Rates Assessment Answers Public Domain eBooks
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 - Reaction Rates Assessment Answers Budget-Friendly Options
6. Navigating Reaction Rates Assessment Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Reaction Rates Assessment Answers Compatibility with Devices
 - Reaction Rates Assessment Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reaction Rates Assessment Answers
 - Highlighting and Note-Taking Reaction Rates Assessment Answers
 - Interactive Elements Reaction Rates Assessment Answers
8. Staying Engaged with Reaction Rates Assessment Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reaction Rates Assessment Answers
9. Balancing eBooks and Physical Books Reaction Rates Assessment Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reaction Rates Assessment Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Reaction Rates Assessment Answers
 - Setting Reading Goals Reaction Rates Assessment Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Reaction Rates Assessment Answers
 - Fact-Checking eBook Content of Reaction Rates Assessment Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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