

KEY

Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium Chapters 17 & 18

Essential Skills/ State Standards:

- 9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ($q = m \cdot c \cdot \Delta T$) & latent heat of phase change ($q = m \cdot \Delta H$). (CSS: 7d)
 9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d*, 8c)
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure. (CSS: 9a)
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

Vocabulary

- | | | | |
|----------------|-------------------|--|---------------------------|
| • Condensation | • Deposition | • Calorimeter | • Heat of reaction |
| • Freezing | • Heat | • Joule | • Le Chatlier's principle |
| • Vaporization | • Temperature | • Specific heat | • Equilibrium |
| • Melting | • Thermochemistry | • $q = m \cdot c \cdot \Delta T$ | • Activation Energy |
| • Boiling | • Exothermic | • $q = m \cdot \Delta H_{\text{vaporization}}$ | • Catalyst |
| • Sublimation | • Endothermic | • $q = m \cdot \Delta H_{\text{fusion}}$ | |

Study Guide (Use notes, worksheets, and labs in addition to this study guide)

1. According to the collision theory, what 3 things must happen in order for two molecules to react?

Must collide with each other, Must collide with enough energy, & in the correct orientation

2. What is the importance of the activation energy in a chemical reaction?

Amount of energy needed for a reaction to occur.

3. Using the collision theory, explain if the following would increase/ decrease the rate of reaction:

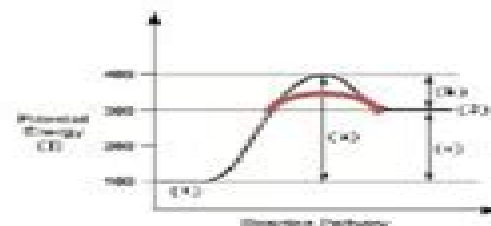


- Increase temp: **Increase, molecules collide more**
- Increase HCl: **Increase, more collisions because more reactants to collide**
- Increase pressure: **Hard to say, wouldn't affect the reactants directly because none of them are gases. One of the products is a gas, so it might cause the gas particles to collide more with the reactants and could slow it down. Tough call.. (this one won't be on the test).**
- Remove MgCl_2 : **Increase, less products means less chance for them to interfere with reactant collisions**
- Add H_2 : **Decrease, more products mean they interfere more with reactants and the reactants would collide less.**

4. Identify the following terms with their correct letter from the graphs:

- Reactants- **1**
- Products- **2**
- Activation energy (number/ letter & amount)-
A (300 Joules)
- Is this reaction exothermic or endothermic?
Endothermic

5. Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do). It lowers the activation energy (see on graph).
Speeds it up by lowering the activation energy.



Reaction Rates And Equilibrium Test And Answers

Disha Experts



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Solutions of Reaction Rate Equations Pertaining to Electron Irradiation of 4:1 Mixtures of N₂ and O₂ Frankline E. Niles, Edna L. Lortie, 1967 One of the Keneshea computer codes see AD 424 173 was adapted for use on the Ballistic Research Laboratories Electronic Scientific Computer Using this modified code reaction rate equations were solved for the following 15 species e NO₂ O O₂ O₃ N₂ NO O O₂ N NO N₂O NO₂ O and O₃ The calculations were made for a 4 1 mixture of N₂ and O₂ at 1 torr total pressure and 300K Rate constants as given by Keneshea and Fowler see AD 646 975 were used The solutions are presented as number densities versus time after the start of the irradiating electron beam A description of the modified code is presented Author **JEE Main 2020 January & September Attempt Solved Papers by Career Point Kota** Career Point Kota, 2020-12-01 Whenever a student decides to prepare for any examination her his first and foremost curiosity is about the type of questions that he she has to face Keeping this in mind we present before you this book containing date wise and shift wise all solved papers of JEE Main 2020 January September Attempt with answer and solutions to majority of questions Solutions to the questions are not just sketch rather have been written in such a manner that the students will be able to understand the application of concept and can answer some other related questions too Salient features of the book are Covers all papers of JEE Main 2020 January September attempt Detailed Errorless Solutions for self evaluation We firmly believe that the book in this form will definitely help a genuine hardworking student for Upcoming Attempt of JEE Main 2021 We have tried our best to keep errors out of this book Comment and criticism from readers will be highly appreciated and incorporated in the subsequent edition We wish to utilize the opportunity to place on record our special thanks to all team members of Content Development for their efforts to make this wonderful book Career Point Ltd

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