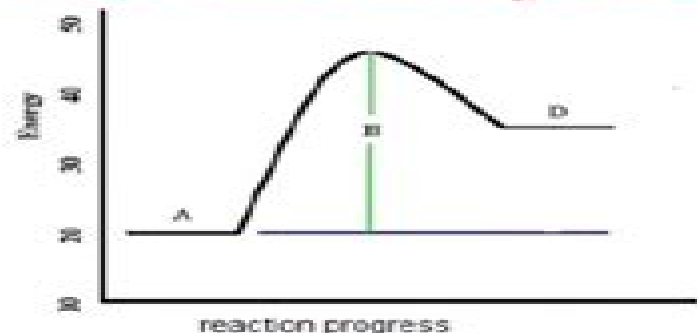


- Suppose two molecules that can react collide. What **3** circumstances need to occur for the molecules to react?
1. There must be enough energy, 2. They must collide, 3. They must have the correct orientation
- What does the **activation energy** for a chemical reaction mean?
This is the amount of energy required to make the molecules collide so that a reaction may occur.
- What role does the **reactivity of the reactants** play in determining the rate of a chemical reaction?
If reactants are not reactive, the reaction may not occur.
- Using the **collision theory**, explain why sugar sprinkled over a Bunsen burner reacts more quickly than a **chunk** of the same solid.
When the sugar is broken up into individual crystals, there is more surface area and more molecules available to collide and react.
- Use the **collision theory** to explain why **increasing the concentration** of a reactant usually increases the reaction rate.
Increasing the concentration increases the number of molecules which increases the likelihood that collisions will occur.
- Use the **collision theory** to explain why **increasing the temperature** usually increases the reaction rate.
Increasing the temperature increases the speed of the molecules. Faster moving molecules collide more.
- Explain how a **catalyst** affects the activation energy for a chemical reaction.
A catalyst ensure that molecules will collide in the correct orientation, so less energy is needed.
- On** the accompanying energy diagram, label the following items:
 - reactants **A**
 - products **D**
 - activation energy **B**
- Draw and label** what this diagram would look like if a **catalyst** was added to the reaction.
- Graph reading
 - How much energy do the reactants have?
20 kJ
 - How much energy do the products have?
35 kJ
 - How much energy is required to activate this un-catalyzed reaction?
Approximately 25 kJ.
 - Will this reaction need to absorb energy or will it release energy to the environment? How do you know?
It will absorb energy because the products have more energy than the reactants.



Reaction Rates Worksheet With Answers

Loleen Berdahl, Jonathan Malloy



Reaction Rates Worksheet With Answers:

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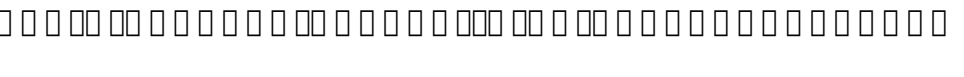
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