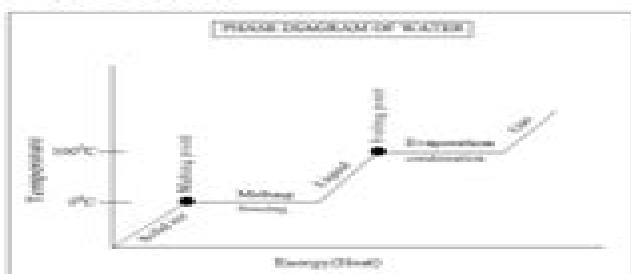


Topic 1: Stoichiometry

1.1 Introduction to Matter & Chemical Change

- **matter:** has mass, consumes volume, made up of particles in constant motion
- **atomic theory:** all matter is composed of atoms that cannot be created nor destroyed
- **states of matter:**
 - **solid:** fixed volume/shape; cannot be compressed; attractive forces hold particles in close-packed arrangement; particles vibrate in fixed positions
 - **liquid:** fixed volume, no fixed shape; cannot be compressed; weaker forces; particles move around
 - **gas:** no fixed volume/shape; can be compressed
- as temperature ↑, KE_{ave} ↑
 - Kelvin (K): SI unit of temperature; $^{\circ}\text{C} + 273$
- **changes of state:**



- **elements:** atoms of one type → form compounds
- **pure substance:** matter with constant composition → proportions are distinct
 - combine to form mixtures
 - homogeneous mixture: uniform composition and proportions
 - heterogeneous mixture: non-uniform composition and varying proportions
- **ion:** charged species
 - **anions:** negatively charged (contain excess electrons)
 - **cations:** positively charged (contains fewer electrons)
- **types of reactions:**
 - combination/synthesis rxns: combination of multiple substances
 - decomposition: breaking down of a single reactant
 - single displacement: one element replaces another in a compound
 - double displacement: swapping of elements

1.2 The Mole Concept

- **mole:** number of particles in a substance
 - fixed amount → Avogadro's number: 6.02×10^{23}
- **isotopes:** different number of neutrons/mass number; same number of protons
 - relative abundance: measure of the percentage of isotopes in a sample
 - relative atomic mass (A_r): weighted average of atomic masses of isotopes/relative abundance
 - relative molecular mass (M_r): combination of all A_r s
- **molar mass:** mass of 1 mol of substance (g/mol)

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