

Question	Answer	Marks	Guidance
5 (a)	<p><math>F_2</math> forces mark  <math>F_2</math> has van der Waals' (forces)  <b>OR</b>  <math>F_2</math> has induced dipole attractions <b>OR</b> interactions  <b>OR</b>  <math>F_2</math> has temporary <b>OR</b> instantaneous dipole(-dipole) attraction <b>OR</b> interactions ✓</p> <p>HCl forces mark  HCl has <b>permanent</b> dipole(-dipole) attractions <b>OR</b> interactions ✓</p> <p>Comparison of strength of forces between molecules mark  intermolecular force in HCl is stronger than that in <math>F_2</math>  <b>OR</b>  permanent dipoles are stronger (than induced dipoles) ✓</p> <p>Boiling point mark  more energy is required to break stronger (intermolecular) forces ✓</p>	4	<p>Use annotations ie ticks crosses ECF ^ etc for this part</p> <p><b>ALLOW</b> vdWs for van der Waals'  <b>IGNORE</b> <math>F_2</math> has covalent bond for this mark  <b>IGNORE</b> <math>F_2</math> has 'intermolecular forces'</p> <p>Quality of written communication: 'dipole(s)' spelled correctly and used in context for the <b>second</b> marking point  <b>IGNORE</b> HCl has 'intermolecular forces'  <b>IGNORE</b> van der Waals' forces in HCl/  <b>DO NOT ALLOW</b> hydrogen bonding  <b>DO NOT ALLOW</b> ionic bonding</p> <p>Look for strength of force comparison anywhere in the answer  <b>ALLOW</b> ECF for hydrogen bonding in HCl being stronger than the stated intermolecular forces in <math>F_2</math>  <b>BUT DO NOT ALLOW</b> this mark if HCl or <math>F_2</math> has covalent bonds broken <b>OR</b> if HCl has ionic bonds broken (the question asks for forces between molecules)  <b>IGNORE</b> HCl has stronger van der Waals' (forces) than <math>F_2</math> (as they both have the same number of electrons)</p> <p><b>DO NOT ALLOW</b> fourth mark if covalent bonds are broken in HCl or <math>F_2</math> <b>OR</b> if ionic bonds are broken in HCl</p> <p><b>IGNORE</b> 'heat' but <b>ALLOW</b> 'heat energy'</p>

# Ocr F321 Jan 2013 Paper

**SJ Ball**



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## **Table of Contents Ocr F321 Jan 2013 Paper**

1. Understanding the eBook Ocr F321 Jan 2013 Paper
  - The Rise of Digital Reading Ocr F321 Jan 2013 Paper
  - Advantages of eBooks Over Traditional Books
2. Identifying Ocr F321 Jan 2013 Paper
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ocr F321 Jan 2013 Paper
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr F321 Jan 2013 Paper
  - Personalized Recommendations
  - Ocr F321 Jan 2013 Paper User Reviews and Ratings
  - Ocr F321 Jan 2013 Paper and Bestseller Lists
5. Accessing Ocr F321 Jan 2013 Paper Free and Paid eBooks
  - Ocr F321 Jan 2013 Paper Public Domain eBooks
  - Ocr F321 Jan 2013 Paper eBook Subscription Services
  - Ocr F321 Jan 2013 Paper Budget-Friendly Options

6. Navigating Ocr F321 Jan 2013 Paper eBook Formats
  - ePub, PDF, MOBI, and More
  - Ocr F321 Jan 2013 Paper Compatibility with Devices
  - Ocr F321 Jan 2013 Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ocr F321 Jan 2013 Paper
  - Highlighting and Note-Taking Ocr F321 Jan 2013 Paper
  - Interactive Elements Ocr F321 Jan 2013 Paper
8. Staying Engaged with Ocr F321 Jan 2013 Paper
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ocr F321 Jan 2013 Paper
9. Balancing eBooks and Physical Books Ocr F321 Jan 2013 Paper
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ocr F321 Jan 2013 Paper
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ocr F321 Jan 2013 Paper
  - Setting Reading Goals Ocr F321 Jan 2013 Paper
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr F321 Jan 2013 Paper
  - Fact-Checking eBook Content of Ocr F321 Jan 2013 Paper
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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