

- (b) A solution is prepared by the addition of 10.0 mL of 0.100 M  $\text{CH}_3\text{COOH}$  and 10.0 mL of 0.100 M  $\text{NaOH}$ .

- (c) The solution is a buffer solution of  $\text{CH}_3\text{COOH}$  and  $\text{CH}_3\text{COO}^-$ .

The total volume is 20.0 mL. The concentration of  $\text{CH}_3\text{COOH}$  is 0.050 M.

In this case, we should consider both of the species involved that could be acting as a buffer:  $\text{CH}_3\text{COOH}$  and  $\text{CH}_3\text{COO}^-$ .

(d) The  $\text{CH}_3\text{COOH}$  and  $\text{CH}_3\text{COO}^-$  are 0.050 M and 0.050 M.

$$[\text{CH}_3\text{COOH}] = 0.050 \text{ M} \quad [\text{CH}_3\text{COO}^-] = 0.050 \text{ M}$$

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \quad [\text{CH}_3\text{COO}^-] = [\text{CH}_3\text{COOH}] = 0.050 \text{ M}$$

$$K_a = \frac{[\text{H}^+]}{0.050 \text{ M}} \quad [\text{H}^+] = K_a \times 0.050 \text{ M} = 1.8 \times 10^{-5} \text{ M}$$

$$\text{pH} = -\log[\text{H}^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

$$\text{pH} = 4.74$$

Reaction:  $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}_3\text{O}^+$   
Initial: 0.100 M  $\text{CH}_3\text{COOH}$ , 0.100 M  $\text{NaOH}$   
Change: -0.100 M  $\text{CH}_3\text{COOH}$ , +0.100 M  $\text{CH}_3\text{COO}^-$   
Equilibrium: 0.000 M  $\text{CH}_3\text{COOH}$ , 0.200 M  $\text{CH}_3\text{COO}^-$   
The solution is a buffer. The pH is 4.74.

- (b) The solution is a buffer solution of  $\text{CH}_3\text{COOH}$  and  $\text{CH}_3\text{COO}^-$ . The concentration of  $\text{CH}_3\text{COOH}$  is 0.050 M and the concentration of  $\text{CH}_3\text{COO}^-$  is 0.050 M.

In this case, we should consider both of the species involved that could be acting as a buffer:  $\text{CH}_3\text{COOH}$  and  $\text{CH}_3\text{COO}^-$ .



- (c) The solution is a buffer solution of  $\text{CH}_3\text{COOH}$  and  $\text{CH}_3\text{COO}^-$ . The concentration of  $\text{CH}_3\text{COOH}$  is 0.050 M and the concentration of  $\text{CH}_3\text{COO}^-$  is 0.050 M.

In this case, we should consider both of the species involved that could be acting as a buffer:  $\text{CH}_3\text{COOH}$  and  $\text{CH}_3\text{COO}^-$ .

- (d) The  $\text{CH}_3\text{COOH}$  and  $\text{CH}_3\text{COO}^-$  are 0.050 M and 0.050 M.

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \quad [\text{CH}_3\text{COO}^-] = [\text{CH}_3\text{COOH}] = 0.050 \text{ M}$$

$$K_a = \frac{[\text{H}^+]}{0.050 \text{ M}} \quad [\text{H}^+] = K_a \times 0.050 \text{ M} = 1.8 \times 10^{-5} \text{ M}$$

$$\text{pH} = -\log[\text{H}^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

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**Richard Bailey**



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