

## Pre-Lab Information

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|                     |                                                                                                                                                                                                                                                                                                                               |
|---------------------|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| <b>Purpose</b>      | Explore the rate of a chemical reaction at different temperatures and with different reactant particle sizes.                                                                                                                                                                                                                 |
| <b>Time</b>         | Approximately 45 minutes                                                                                                                                                                                                                                                                                                      |
| <b>Question</b>     | What are the effects of temperature and a reactant's particle size on reaction rate?                                                                                                                                                                                                                                          |
| <b>Hypothesis 1</b> | If you increase the temperature of a reaction, then the reaction rate will increase because particles experience more collisions at higher temperatures.                                                                                                                                                                      |
| <b>Hypothesis 2</b> | If you decrease the particle size of a reactant, then the reaction rate will increase because more of the reactants' surface area is exposed allowing more particles to make contact with each other.                                                                                                                         |
| <b>Variables</b>    | Independent variables: particle size, temperature<br>Dependent variable: reaction rate                                                                                                                                                                                                                                        |
| <b>Summary</b>      | The reaction between effervescent tablets and water will produce $\text{CO}_2$ gas, which you can observe in the form of bubbles; when the bubbling stops, the reaction is complete. You will measure the reaction times and rates for tablets dissolved at four temperatures and broken into three different particle sizes. |

## Safety

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- Always wear a lab coat and safety goggles when performing an experiment.
- Behavior in the lab needs to be purposeful. Use caution when using the hot plate, as it can cause burns.
- Check glassware such as graduated cylinders and beakers for cracks and chips prior to use.
- Use the right gear, such as chemically resistant gloves and oven mitts, when performing the experiment.
- Report all accidents—no matter how big or small—to your teacher.

## Procedure

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### Step 1: Gather Materials

- 250 mL graduated cylinder
- Thermometer
- Water
- Timer
- Four 250 mL beakers
- Seven 1,000 mg effervescent tablets
- Two pieces of filter paper
- 600 mL beaker
- Ice
- Hot plate

# Reaction Rates Chemical Study Guide

**R Bogdan**



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