#### Worksheet- Reaction Rates

Use this reaction for the questions below:  $C_0H_{22}O_0$  (s) + 6  $O_2(g) \rightarrow$  6  $H_2O$  (g) + 6  $CO_2$  (g)

- 1. What happens to the concentrations of:
  - a. C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> & O<sub>2</sub> as the reaction proceeds →?
- b. H<sub>2</sub>O + CO<sub>2</sub> as the reaction proceeds →?
- 2. According to the collision theory, what 3 circumstances are needed for CoHizOo & Oz to react?
- 3. What is the activation energy for a chemical reaction?
- 4. Use the equation & the collision theory to explain:

$$C_6H_{12}O_6(s) + 6 O_2(g) \rightarrow 6 H_2O(g) + 6 CO_2(g)$$

	Change in condition:	Does this <u>increase</u> or <u>decrease</u> the rate of reaction?	Explain why
a.	Increasing the temperature	Exi (normanca (apenda up)	Ext Molecules more faster to collide more - Increased on rate
ь.	Increasing the concentration of CoHorOs		
C.	Decreasing the concentration of O <sub>2</sub>		
d.	Increase the surface area by chewing up food in your mouth		
6.	Decreasing the temperature		
f.	Increasing the pressure in the container		
9	Decreasing the concentration of H <sub>2</sub> O		
h.	Increasing the volume of the container the reaction occurs in		
	Increasing the concentration of CO <sub>2</sub>		
1	Using a catalyst (like solivary amylase)		

- On the accompanying energy diagram, <u>label</u> the following terms:
  - a. reactants b. products c. activation energy
- On the graph to the right, <u>draw</u> and <u>label</u> what this diagram would look like if a <u>catalyst</u> was added to the reaction.
- 7. Graph reading
  - a. How much energy (#) do the reactants have?
  - b. How much energy (#) do the products have?
  - c. How much energy (#) is required to activate this un-catalyzed reaction?
  - d. Is this reaction endothermic or exothermic? How do you know?



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