

- Q6 When water was added to the anhydrate of $\text{CuSO}_4(x)$, did the reaction proceed in the direction of the reactants or products?

B.2 Reaction of Copper(II) Ion and Hydroxide

	Test tube 1	Test tube 2
Initial color of CuCl_2 solution (1)	light blue	light blue
Initial appearance $\text{Cu(OH)}_2(x)$ (2)	clear	clear
Change in appearance (3)	cloudy	clear
Change in appearance (4)	cloudy	clear

- Q7 In test tube 1, did the addition of OH^- cause the reaction to go in the direction of the reactants or products? Explain.
- Q8 In test tube 2, did the addition of H_3O^+ cause the reaction to go in the direction of the reactants or products? Explain.

C. Changing Equilibrium Conditions: Le Châtelier's Principle

C.1 Effect of Changing Reactant Concentration

Test tube	Color	Is color deeper or lighter red compared to the reference?	$[\text{FeSCN}^{2+}]$ increases or decreases?
1. Reference	red-orange	deeper/lighter	increases/decreases
2. Fe^{3+} added	dark red	deeper	increased
3. SCN^- added	blood red	deeper	increased
4. Fe^{3+} removed	orange	lighter	decreased

Reaction Rates And Equilibrium Lab Answers

BM King



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Reaction Rates for High-temperature Air with Carbon and Sodium Impurities Mina L. Carnicom, 1968 The values used by a number of investigators for the rate constants of high temperature greater than or equal to 1000 K homogeneous gaseous reactions involving species of the elements nitrogen oxygen carbon and sodium have been compiled and are presented in tabular form Included are reactions involving neutral species charged species free electrons some species in excited electronic or vibrational states and radiative processes Safety-Scale Lab Exp Biochem 2e Spencer L.

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