

Name: _____

Date: _____

Student Exploration: Reaction Energy

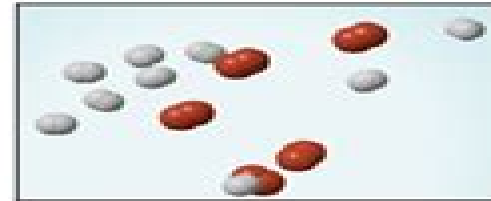
Vocabulary: calorimeter, chemical bond, endothermic, enthalpy, exothermic, Hess's law

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. Two magnets are stuck together. What might you have to do to get them to separate?
Pull them apart
2. Suppose you held two magnets a short distance apart, then let go. What would happen?
They would connect
3. Think about the magnets in terms of energy. In which case do you increase the potential energy of the magnets? In which case do you increase the kinetic energy of the magnets?

Gizmo Warm-up

Just like magnets, atoms of different elements are attracted together to form **chemical bonds**. Breaking these bonds requires energy. When a new bond forms, energy is released and temperatures rise. In the *Reaction Energy Gizmo™*, you will explore how the energy of chemical bonding relates to temperature changes that occur during chemical reactions.



To begin, check that **Reaction 1** and **Forward** are selected. In this reaction, hydrogen (H_2) and oxygen (O_2) react to form water (H_2O). The reaction takes place inside a device called a **calorimeter**. Inside the calorimeter, a small chamber holds the reactants. The rest of the calorimeter is filled with water.

1. Click **Play** (▶). What happens? When I clicked play, all of the separate H_2 and O_2 molecules formed to create H_2O .
2. How does the temperature change? The temperature is increased while the molecules are reacting to form water, and the increase in temperature slows once the molecules are all bonded to create the H_2O molecules. The temperature raises from 21.0 Degrees Celsius to 26.0 Degrees Celsius while the molecules were bonding.

Reaction Energy Study Guide Answers

Didier Musso



Reaction Energy Study Guide Answers:

"O" Level Study Guide - Chemistry Quite Easily Done , Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section s critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text

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Metz, 1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences **Chemical**

Principles Study Guide/Solutions Manual John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications *Study Guide with Answers to Selected Problems* H. Stephen Stoker, Danny V.

White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises *Chemical*

Principles Student's Study Guide & Solutions Manual John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text Study Guide and Partial Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry Susan

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