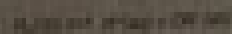


16. The dissociation of water is shown below:



At 25 °C, the ion product of water,  $K_w$ , is  $1.0 \times 10^{-14}$ .

At 25 °C, the ion product of water,  $K_w$ , is  $1.0 \times 10^{-14}$ .

17. Explain what the dissociation of water is an indication of water's behavior.

Endothermic

18. Predict, using a calculation, whether a solution of 1.0 M  $HCl$  is acidic, basic, or neutral. (2)

Acidic

Explain by calculating pH.

19.  $pK_a$ ,  $pH$ , and  $pOH$  are significant values.

Calculate  $pK_a$  of  $HCl$ .

Give your answer to two decimal places.

$$-\log(4.31 \times 10^{-4}) = 3.07$$

$$pK_a = 3.07$$

20. A solution of hydrochloric acid is shown with water and the resulting ions in solution at 25 °C.

Calculate the pH of the resulting solution at 25 °C.

Give your answer to two decimal places.

$$pH = 1.30$$

pH = \_\_\_\_\_ (2)

(Total: 10)

# Ocr Chemistry A2 Unofficial Mark Scheme 2014

**Christian G. Meyer**



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## **Table of Contents Ocr Chemistry A2 Unofficial Mark Scheme 2014**

1. Understanding the eBook Ocr Chemistry A2 Unofficial Mark Scheme 2014
  - The Rise of Digital Reading Ocr Chemistry A2 Unofficial Mark Scheme 2014
  - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry A2 Unofficial Mark Scheme 2014
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ocr Chemistry A2 Unofficial Mark Scheme 2014
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry A2 Unofficial Mark Scheme 2014
  - Personalized Recommendations
  - Ocr Chemistry A2 Unofficial Mark Scheme 2014 User Reviews and Ratings
  - Ocr Chemistry A2 Unofficial Mark Scheme 2014 and Bestseller Lists
5. Accessing Ocr Chemistry A2 Unofficial Mark Scheme 2014 Free and Paid eBooks
  - Ocr Chemistry A2 Unofficial Mark Scheme 2014 Public Domain eBooks
  - Ocr Chemistry A2 Unofficial Mark Scheme 2014 eBook Subscription Services
  - Ocr Chemistry A2 Unofficial Mark Scheme 2014 Budget-Friendly Options
6. Navigating Ocr Chemistry A2 Unofficial Mark Scheme 2014 eBook Formats
  - ePub, PDF, MOBI, and More
  - Ocr Chemistry A2 Unofficial Mark Scheme 2014 Compatibility with Devices
  - Ocr Chemistry A2 Unofficial Mark Scheme 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ocr Chemistry A2 Unofficial Mark Scheme 2014
  - Highlighting and Note-Taking Ocr Chemistry A2 Unofficial Mark Scheme 2014
  - Interactive Elements Ocr Chemistry A2 Unofficial Mark Scheme 2014
8. Staying Engaged with Ocr Chemistry A2 Unofficial Mark Scheme 2014

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ocr Chemistry A2 Unofficial Mark Scheme 2014
9. Balancing eBooks and Physical Books Ocr Chemistry A2 Unofficial Mark Scheme 2014
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ocr Chemistry A2 Unofficial Mark Scheme 2014
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ocr Chemistry A2 Unofficial Mark Scheme 2014
- Setting Reading Goals Ocr Chemistry A2 Unofficial Mark Scheme 2014
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr Chemistry A2 Unofficial Mark Scheme 2014
- Fact-Checking eBook Content of Ocr Chemistry A2 Unofficial Mark Scheme 2014
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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