

**QUESTION 3**

3. 1.Learners made a mini volcano in science laboratory by adding sodium bicarbonate to ethanoic acid. They added 100 ml of a  $0,2 \text{ mol.dm}^{-3}$  ethanoic acid solution to 10 g of  $\text{NaHCO}_3$  to start the reaction of the volcano.

The balanced equation for this reaction is:



- 3.1.1 Define the term limiting reagent. (2)
- 3.1.2 Determine the limiting reagent in this reaction. (6)
- 3.1.3 Calculate the mass of the other substance in excess. (3)
- 3.1.4 Calculate the volume of  $\text{CO}_2$  produced at STP. (4)

3.2 A solution of potassium hydroxide (KOH) is prepared by dissolving 3,36 g crystals of KOH in  $250 \text{ cm}^3$  of water.

Calculate the concentration of the potassium hydroxide solution. (4)

3.3  $25 \text{ cm}^3$  of a potassium hydroxide solution of concentration  $0,25 \text{ mol.dm}^{-3}$  completely neutralises a dilute solution of sulphuric acid ( $\text{H}_2\text{SO}_4$ ) in a flask.

The unbalanced equation below represents the reaction that takes place:



- 3.3.1 Balance the above equation. (2)
- 3.3.2 Calculate the mass of sulphuric acid in the flask. (5)

**[25]**

# Physical Science Grade 11 Controlled Test 2 September 2014

**Rachel Sandford**



**Physical Science Grade 11 Controlled Test 2 September 2014:**

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## **Table of Contents Physical Science Grade 11 Controlled Test 2 September 2014**

1. Understanding the eBook Physical Science Grade 11 Controlled Test 2 September 2014
  - The Rise of Digital Reading Physical Science Grade 11 Controlled Test 2 September 2014
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Grade 11 Controlled Test 2 September 2014
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science Grade 11 Controlled Test 2 September 2014
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Grade 11 Controlled Test 2 September 2014
  - Personalized Recommendations
  - Physical Science Grade 11 Controlled Test 2 September 2014 User Reviews and Ratings
  - Physical Science Grade 11 Controlled Test 2 September 2014 and Bestseller Lists
5. Accessing Physical Science Grade 11 Controlled Test 2 September 2014 Free and Paid eBooks
  - Physical Science Grade 11 Controlled Test 2 September 2014 Public Domain eBooks
  - Physical Science Grade 11 Controlled Test 2 September 2014 eBook Subscription Services
  - Physical Science Grade 11 Controlled Test 2 September 2014 Budget-Friendly Options
6. Navigating Physical Science Grade 11 Controlled Test 2 September 2014 eBook Formats
  - ePub, PDF, MOBI, and More
  - Physical Science Grade 11 Controlled Test 2 September 2014 Compatibility with Devices
  - Physical Science Grade 11 Controlled Test 2 September 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Science Grade 11 Controlled Test 2 September 2014
  - Highlighting and Note-Taking Physical Science Grade 11 Controlled Test 2 September 2014
  - Interactive Elements Physical Science Grade 11 Controlled Test 2 September 2014
8. Staying Engaged with Physical Science Grade 11 Controlled Test 2 September 2014

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Physical Science Grade 11 Controlled Test 2 September 2014
- 9. Balancing eBooks and Physical Books Physical Science Grade 11 Controlled Test 2 September 2014
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science Grade 11 Controlled Test 2 September 2014
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Grade 11 Controlled Test 2 September 2014
  - Setting Reading Goals Physical Science Grade 11 Controlled Test 2 September 2014
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Grade 11 Controlled Test 2 September 2014
  - Fact-Checking eBook Content of Physical Science Grade 11 Controlled Test 2 September 2014
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
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  - Interactive and Gamified eBooks

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