

QUESTION 10

The three test tubes, X, Y and Z below, contain colourless solutions. The solutions in each test tube can either be potassium iodide or sodium carbonate or magnesium sulphate. Tests are carried out to determine which test tube contains which salts.



The following tests were carried out and the observations were made:

- Barium chloride solution is added to solution X; the solution remains colourless.
- Barium chloride solution is added to Y and Z; a white precipitate is formed in both cases.
- Dilute nitric acid is added to the precipitate in solution Y; the precipitate dissolves.

(2)

10.1 Define the term *dissociation*.

10.2 Give a reason why barium chloride solution can be used as electrolyte.

(2)

10.3 Write down the name of the precipitate in test tube Z.

(1)

Use the above information to identify the solutions in each of the test tubes X, Y and Z.

10.4 Write down the LETTER that represents the test tube and next to it, the NAME of the solution and the reason for the choice.

(6)

10.5 What type of reaction takes place between the precipitate in Y and the nitric acid?

(1)

[12]

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1. Understanding the eBook Physical Science Exemplar 2014 November Grade 10
 - The Rise of Digital Reading Physical Science Exemplar 2014 November Grade 10
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Exemplar 2014 November Grade 10
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Exemplar 2014 November Grade 10
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Exemplar 2014 November Grade 10
 - Personalized Recommendations
 - Physical Science Exemplar 2014 November Grade 10 User Reviews and Ratings
 - Physical Science Exemplar 2014 November Grade 10 and Bestseller Lists
5. Accessing Physical Science Exemplar 2014 November Grade 10 Free and Paid eBooks

- Physical Science Exemplar 2014 November Grade 10 Public Domain eBooks
 - Physical Science Exemplar 2014 November Grade 10 eBook Subscription Services
 - Physical Science Exemplar 2014 November Grade 10 Budget-Friendly Options
6. Navigating Physical Science Exemplar 2014 November Grade 10 eBook Formats
- ePub, PDF, MOBI, and More
 - Physical Science Exemplar 2014 November Grade 10 Compatibility with Devices
 - Physical Science Exemplar 2014 November Grade 10 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Physical Science Exemplar 2014 November Grade 10
 - Highlighting and Note-Taking Physical Science Exemplar 2014 November Grade 10
 - Interactive Elements Physical Science Exemplar 2014 November Grade 10
8. Staying Engaged with Physical Science Exemplar 2014 November Grade 10
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Exemplar 2014 November Grade 10
9. Balancing eBooks and Physical Books Physical Science Exemplar 2014 November Grade 10
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Exemplar 2014 November Grade 10
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Exemplar 2014 November Grade 10
- Setting Reading Goals Physical Science Exemplar 2014 November Grade 10
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Exemplar 2014 November Grade 10
- Fact-Checking eBook Content of Physical Science Exemplar 2014 November Grade 10
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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