Chapter 9

9.1 Compounds and Chemical Change

- Atom- the smallest unit of an element that can exist alone or in combination with other elements.
- Compounds- formed when atoms are held together by an attractive force called a chemical bond
- Molecule- The smallest particle of a compound or a gaseous element, that can exist and still remain the characteristic chemical properties of a substance
- Chemical reaction

 a change in matter where different chemical substances are created by breaking and/or forming chemical bonds.
 - Examples: Photosynthesis
- Energy is absorbed -> produce new chemical substances -> more internal potential energy
- Energy is released when new chemical substances are produced -> less internal potential energy
- Chemical Equation: a way of describing what happens during a chemical reaction
 - Example: Photosynthesis equation-

Energy/Sunlight + carbon dioxide molecules + water molecules -> plant material molecules + oxygen molecules

9.2 Valence Electrons and Ions

- the number of electrons in the outermost orbital that usually determines the chemical properties of an atom
- <u>Valence electrons</u> the number of electrons in the outermost orbital that usually determines the properties of an atom
- Octet Rule after the element helium, atoms attempt to acquire an outer orbital with eight electrons through chemical reactions

9.3 Chemical Bonds

- chemical bond-an attractive force that holds atoms together in a compound.
- Three Classes of Chemical Bonds-
 - Ionic Bonds
 - Covalent Bonds
 - Metallic Bonds
- Fonic Bonds
 - type of linkage formed from the electrostatic attraction between oppositely
 charged ions in a chemical compound
- Covalent bonds: result when atoms achieve the noble gas electron structure by sharing
 electrons. They generally formed between the nonmetallic elements on the right side of
 the periodic table
- Metallic bonds: formed in solid metals such as iron, copper, and the other metallic elements that make up about 80 percent of all the elements.
- . Heat formation- energy that is released during an ionic bond formation
- Boemula- describes what elements are in the compound and in what proportions
- Simple bond covalent bond
- Double bond- formed by two pairs of electrons that are shared by two atoms

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