Chapter 9

9.1 Compounds and Chemical Change

- Atom- the smallest unit of an element that can exist alone or in combination with other elements.
- Compounds- formed when atoms are held together by an attractive force called a chemical bond
- Molecule- The smallest particle of a compound or a gaseous element, that can exist and still remain the characteristic chemical properties of a substance
- Chemical reaction

 a change in matter where different chemical substances are created by breaking and/or forming chemical bonds.
 - Examples: Photosynthesis
- Energy is absorbed -> produce new chemical substances -> more internal potential energy
- Energy is released when new chemical substances are produced -> less internal potential energy
- Chemical Equation: a way of describing what happens during a chemical reaction
 - Example: Photosynthesis equation-

Energy/Sunlight + carbon dioxide molecules + water molecules -> plant material molecules + oxygen molecules

9.2 Valence Electrons and Ions

- the number of electrons in the outermost orbital that usually determines the chemical properties of an atom
- <u>Valence electrons</u> the number of electrons in the outermost orbital that usually determines the properties of an atom
- Octet Rule after the element helium, atoms attempt to acquire an outer orbital with eight electrons through chemical reactions

9.3 Chemical Bonds

- chemical bond-an attractive force that holds atoms together in a compound.
- Three Classes of Chemical Bonds-
 - Ionic Bonds
 - Covalent Bonds
 - Metallic Bonds
- Fonic Bonds
 - type of linkage formed from the electrostatic attraction between oppositely charged ions in a chemical compound
- Covalent bonds: result when atoms achieve the noble gas electron structure by sharing
 electrons. They generally formed between the nonmetallic elements on the right side of
 the periodic table
- Metallic bonds: formed in solid metals such as iron, copper, and the other metallic elements that make up about 80 percent of all the elements.
- . Heat formation- energy that is released during an ionic bond formation
- Boemula- describes what elements are in the compound and in what proportions
- Single bond covalent bond
- Double bond- formed by two pairs of electrons that are shared by two atoms

Physical Science Chapter Vocabulary

Robert R. Redfield

Physical Science Chapter Vocabulary:

Physical Science Chapter Vocabulary Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has be apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Physical Science Chapter Vocabulary**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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