

Name key

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. When an atom loses an electron it becomes a positive ion (cation), and when it gains an electron it becomes a negative ion (anion).
2. An element in the halogen group has 7 valence electrons, and wants to gain 1 more to have 8. This is because of the octet rule.
3. An ionic bond is the attraction between oppositely charged ions, AKA a metal and a nonmetal.
4. In a covalent bond, atoms share electrons, whereas in an ionic bond, atoms transfer electrons.
5. (Honors only) When electrons are shared unequally in a covalent bond it is polar, and when they're shared equally it is nonpolar. You can predict if a compound is ionic, polar covalent or nonpolar covalent by subtracting the electronegativities.
6. An atom can have up to 8 valence electrons. The family (column) on the periodic table that all have 8 valence electrons is called noble gases.
7. Atoms can gain, lose, or share valence electrons to become more stable.

B. Formulas for Ionic compounds - Fill in the chart below

	Magnesium	Fluorine	Nitrogen	Lithium
# OF VALENCE ELECTRONS	<u>2</u>	<u>7</u>	<u>5</u>	<u>1</u>
OXIDATION #	<u>+2</u>	<u>-1</u>	<u>-3</u>	<u>+1</u>

2. What is the **chemical formula** for each of the following ionic compounds?
(*hint: you can either draw the Lewis structures and use arrows to show the transfer of electrons, OR use the periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.*)

- a) Potassium and Chlorine - KCl
- b) Calcium and Bromine - CaBr₂
- c) Magnesium and oxygen - MgO

Physical Science Chemical Bond Study Guide

Jin-Ying Zhang



Physical Science Chemical Bond Study Guide:

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Council,Division of Behavioral and Social Sciences and Education,Center for Education,Committee on Science Education

K-12,Steering Committee on Taking Stock of the National Science Education Standards: The Research,2002-12-05 In 2001

with support from National Science Foundation the National Research Council began a review of the evidence concerning

whether or not the National Science Education Standards have had an impact on the science education enterprise to date

and if so what that impact has been This publication represents the second phase of a three phase effort by the National

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Education National Research Council 2002 That report provided organizing principles for the design conduct and

interpretation of research regarding the influence of national standards The Framework developed in Phase I was used to

structure the current review of research that is reported here Phase II began in mid 2001 involved a thorough search and

review of the research literature on the influence of the NSES and concludes with this publication which summarizes the

proceedings of a workshop conducted on May 10 2002 in Washington DC Phase III will provide input collected in 2002 from

science educators administrators at all levels and other practitioners and policy makers regarding their views of the NSES

the ways and extent to which the NSES are influencing their work and the systems that support science education and what

next steps are needed **University of South Florida Educational Review** University of South Florida,1966

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