

**EXAMINATION DATA SHEET FOR THE PHYSICAL SCIENCES
(CHEMISTRY)****TABLE 1 PHYSICAL CONSTANTS**

NAME	SYMBOL	VALUE
Magnitude of charge on electron	e	$1,6 \times 10^{-19} \text{ C}$
Mass of an electron	m_e	$9,1 \times 10^{-31} \text{ kg}$
Standard pressure	p^θ	$1,01 \times 10^5 \text{ Pa}$
Molar gas volume at STP	V_m	$22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$
Standard temperature	T^θ	$273 \text{ K (0 } ^\circ\text{C)}$
Avogadro's constant	N_A	$6,02 \times 10^{23} \text{ mol}^{-1}$
Faraday's constant	F	$96\,500 \text{ C} \cdot \text{mol}^{-1}$

TABLE 2 CHEMISTRY FORMULAE

$n = \frac{m}{M}$	$n = \frac{N}{N_A}$	$n = \frac{V}{V_m}$
$c = \frac{n}{V}$ OR $c = \frac{m}{MV}$	$K_w = [\text{H}_3\text{O}^+] \cdot [\text{OH}^-] = 1 \times 10^{-14}$ at $25\, ^\circ\text{C (298 K)}$	
$q = It$ $q = nF$	$E_{\text{cell}}^\theta = E_{\text{cathode}}^\theta - E_{\text{anode}}^\theta$ $E_{\text{cell}}^\theta = E_{\text{oxidising agent}}^\theta - E_{\text{reducing agent}}^\theta$	

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Paper 2014 GS Paper I 10 MPPSC Prelims Solved Previous Paper 2015 GS Paper I 11 MPPSC Prelims Solved Previous Paper 2016 GS Paper I 12 MPPSC Prelims Solved Previous Paper 2017 GS Paper I 13 MPPSC Prelims Solved Previous Paper 2018 GS Paper I 14 MPPSC Prelims Solved Previous Paper 2012 CSAT Paper II 15 MPPSC Prelims Solved Previous Paper 2013 CSAT Paper II 16 MPPSC Prelims Solved Previous Paper 2014 CSAT Paper II 17 MPPSC Prelims Solved Previous Paper 2015 CSAT Paper II 18 MPPSC Prelims Solved Previous Paper 2016 CSAT Paper II 19 MPPSC Prelims Solved Previous Paper 2017 CSAT Paper II 20 MPPSC Prelims Solved Previous Paper 2018 CSAT Paper II

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1. Understanding the eBook Physical Science P2 2014 Prelims
 - The Rise of Digital Reading Physical Science P2 2014 Prelims
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science P2 2014 Prelims
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science P2 2014 Prelims
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science P2 2014 Prelims
 - Personalized Recommendations
 - Physical Science P2 2014 Prelims User Reviews and Ratings
 - Physical Science P2 2014 Prelims and Bestseller Lists
5. Accessing Physical Science P2 2014 Prelims Free and Paid eBooks

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- Physical Science P2 2014 Prelims eBook Subscription Services
- Physical Science P2 2014 Prelims Budget-Friendly Options
- 6. Navigating Physical Science P2 2014 Prelims eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science P2 2014 Prelims Compatibility with Devices
 - Physical Science P2 2014 Prelims Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science P2 2014 Prelims
 - Highlighting and Note-Taking Physical Science P2 2014 Prelims
 - Interactive Elements Physical Science P2 2014 Prelims
- 8. Staying Engaged with Physical Science P2 2014 Prelims
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science P2 2014 Prelims
- 9. Balancing eBooks and Physical Books Physical Science P2 2014 Prelims
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science P2 2014 Prelims
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science P2 2014 Prelims
 - Setting Reading Goals Physical Science P2 2014 Prelims
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science P2 2014 Prelims
 - Fact-Checking eBook Content of Physical Science P2 2014 Prelims
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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