

## Balancing Equations Worksheet

- 1)  $\underline{\hspace{1cm}} \text{Na}_3\text{PO}_4 + \underline{\hspace{1cm}} \text{KOH} \rightarrow \underline{\hspace{1cm}} \text{NaOH} + \underline{\hspace{1cm}} \text{K}_3\text{PO}_4$
- 2)  $\underline{\hspace{1cm}} \text{MgF}_2 + \underline{\hspace{1cm}} \text{Li}_2\text{CO}_3 \rightarrow \underline{\hspace{1cm}} \text{MgCO}_3 + \underline{\hspace{1cm}} \text{LiF}$
- 3)  $\underline{\hspace{1cm}} \text{P}_4 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{P}_2\text{O}_5$
- 4)  $\underline{\hspace{1cm}} \text{RbNO}_3 + \underline{\hspace{1cm}} \text{BeF}_2 \rightarrow \underline{\hspace{1cm}} \text{Be}(\text{NO}_3)_2 + \underline{\hspace{1cm}} \text{RbF}$
- 5)  $\underline{\hspace{1cm}} \text{AgNO}_3 + \underline{\hspace{1cm}} \text{Cu} \rightarrow \underline{\hspace{1cm}} \text{Cu}(\text{NO}_3)_2 + \underline{\hspace{1cm}} \text{Ag}$
- 6)  $\underline{\hspace{1cm}} \text{CF}_4 + \underline{\hspace{1cm}} \text{Br}_2 \rightarrow \underline{\hspace{1cm}} \text{CBr}_4 + \underline{\hspace{1cm}} \text{F}_2$
- 7)  $\underline{\hspace{1cm}} \text{HCN} + \underline{\hspace{1cm}} \text{CuSO}_4 \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{SO}_4 + \underline{\hspace{1cm}} \text{Cu}(\text{CN})_2$
- 8)  $\underline{\hspace{1cm}} \text{GaF}_3 + \underline{\hspace{1cm}} \text{Cs} \rightarrow \underline{\hspace{1cm}} \text{CsF} + \underline{\hspace{1cm}} \text{Ga}$
- 9)  $\underline{\hspace{1cm}} \text{BaS} + \underline{\hspace{1cm}} \text{PtF}_2 \rightarrow \underline{\hspace{1cm}} \text{BaF}_2 + \underline{\hspace{1cm}} \text{PtS}$
- 10)  $\underline{\hspace{1cm}} \text{N}_2 + \underline{\hspace{1cm}} \text{H}_2 \rightarrow \underline{\hspace{1cm}} \text{NH}_3$
- 11)  $\underline{\hspace{1cm}} \text{NaF} + \underline{\hspace{1cm}} \text{Br}_2 \rightarrow \underline{\hspace{1cm}} \text{NaBr} + \underline{\hspace{1cm}} \text{F}_2$
- 12)  $\underline{\hspace{1cm}} \text{Pb}(\text{OH})_2 + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{O} + \underline{\hspace{1cm}} \text{PbCl}_2$
- 13)  $\underline{\hspace{1cm}} \text{AlBr}_3 + \underline{\hspace{1cm}} \text{K}_2\text{SO}_4 \rightarrow \underline{\hspace{1cm}} \text{KBr} + \underline{\hspace{1cm}} \text{Al}_2(\text{SO}_4)_3$
- 14)  $\underline{\hspace{1cm}} \text{CH}_4 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 15)  $\underline{\hspace{1cm}} \text{Na}_3\text{PO}_4 + \underline{\hspace{1cm}} \text{CaCl}_2 \rightarrow \underline{\hspace{1cm}} \text{NaCl} + \underline{\hspace{1cm}} \text{Ca}_3(\text{PO}_4)_2$
- 16)  $\underline{\hspace{1cm}} \text{K} + \underline{\hspace{1cm}} \text{Cl}_2 \rightarrow \underline{\hspace{1cm}} \text{KCl}$
- 17)  $\underline{\hspace{1cm}} \text{Al} + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{H}_2 + \underline{\hspace{1cm}} \text{AlCl}_3$
- 18)  $\underline{\hspace{1cm}} \text{N}_2 + \underline{\hspace{1cm}} \text{F}_2 \rightarrow \underline{\hspace{1cm}} \text{NF}_3$
- 19)  $\underline{\hspace{1cm}} \text{SO}_2 + \underline{\hspace{1cm}} \text{Li}_2\text{Se} \rightarrow \underline{\hspace{1cm}} \text{SSe}_2 + \underline{\hspace{1cm}} \text{Li}_2\text{O}$
- 20)  $\underline{\hspace{1cm}} \text{NH}_3 + \underline{\hspace{1cm}} \text{H}_2\text{SO}_4 \rightarrow \underline{\hspace{1cm}} (\text{NH}_4)_2\text{SO}_4$

# Physical Science Balancing Equations Answer Key

**Robert R. Redfield**



## **Physical Science Balancing Equations Answer Key:**

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**Physical Science** George A. Williams,Richard Barnes,Ray Doerhoff,Max Bolen,1979 A physical science text stressing an awareness of the environment with related laboratory activities to lead the student into discovering basic laws and concepts of physics and chemistry

**Impact of Advances in Computing and Communications Technologies on Chemical Science and Technology** National Research Council,Division on Engineering and Physical Sciences,Commission on Physical Sciences, Mathematics, and Applications,Chemical Sciences Roundtable,1999-10-01 The Chemical Sciences Roundtable provides a forum for discussing chemically related issues affecting government industry and government The goal is to strengthen the chemical sciences by foster communication among all the important stakeholders At a recent Roundtable meeting information technology was identified as an issue of increasing importance to all sectors of the chemical enterprise This book is the result of a workshop

convened to explore this topic      **The Chemical News and Journal of Physical Science** ,1879      *Chemical News and Journal of Physical Science* William Crookes,James H. Gardiner,Gerald Druce,H. W. Blood-Ryan,1916      Chemical News and Journal of Physical Science ,1878      *Physical Chemistry for the Chemical and Biological Sciences* Raymond Chang,2000-05-12 Hailed by advance reviewers as a kinder gentler P Chem text this book meets the needs of an introductory course on physical chemistry and is an ideal choice for courses geared toward pre medical and life sciences students Physical Chemistry for the Chemical and Biological Sciences offers a wealth of applications to biological problems numerous worked examples and around 1000 chapter end problems      **Invariant Manifolds for Physical and Chemical Kinetics** Alexander N. Gorban,Iliya V. Karlin,2005-02-01 By bringing together various ideas and methods for extracting the slow manifolds the authors show that it is possible to establish a more macroscopic description in nonequilibrium systems The book treats slowness as stability A unifying geometrical viewpoint of the thermodynamics of slow and fast motion enables the development of reduction techniques both analytical and numerical Examples considered in the book range from the Boltzmann kinetic equation and hydrodynamics to the Fokker Planck equations of polymer dynamics and models of chemical kinetics describing oxidation reactions Special chapters are devoted to model reduction in classical statistical dynamics natural selection and exact solutions for slow hydrodynamic manifolds The book will be a major reference source for both theoretical and applied model reduction Intended primarily as a postgraduate level text in nonequilibrium kinetics and model reduction it will also be valuable to PhD students and researchers in applied mathematics physics and various fields of engineering      **Holt Science and Technology** Holt Rinehart & Winston,2004-02      **E3 Chemistry Review Book - 2018 Home Edition (Answer Key Included)** Effiong Eyo,2017-10-20 With Answer Key to All Questions Chemistry students and homeschoolers Go beyond just passing Enhance your understanding of chemistry and get higher marks on homework quizzes tests and the regents exam with E3 Chemistry Review Book 2018 With E3 Chemistry Review Book students will get clean clear engaging exciting and easy to understand high school chemistry concepts with emphasis on New York State Regents Chemistry the Physical Setting Easy to read format to help students easily remember key and must know chemistry materials Several example problems with solutions to study and follow Several practice multiple choice and short answer questions at the end of each lesson to test understanding of the materials 12 topics of Regents question sets and 3 most recent Regents exams to practice and prep for any Regents Exam This is the Home Edition of the book Also available in School Edition ISBN 978 197836229 The Home Edition contains an answer key section Teachers who want to recommend our Review Book to their students should recommend the Home Edition Students and and parents whose school is not using the Review Book as instructional material as well as homeschoolers should buy the Home Edition The School Edition does not have answer key in the book A separate answer key booklet is provided to teachers with a class order of the book Whether you are using the school or Home Edition our E3 Chemistry Review Book makes a great supplemental instructional and test prep resource that

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