

# PHYSICAL CHEMISTRY

## (Complete Syllabus)

### SECTION - A

#### Straight Objective Type

This section contains 220 multiple choice questions. Each question has 4 choices (1), (2), (3) and (4), out of which **ONLY ONE** is correct.

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Choose the correct answer :

- The mass of oxygen required to produce enough CO which completely reduces  $x$  gm of  $\text{Fe}_2\text{O}_3$ . Suppose molecular weight of  $\text{O}_2 = y$ , molecular weight of  $\text{Fe}_2\text{O}_3 = z$   
(1)  $\frac{2xz}{3y}$  (2)  $\frac{3xy}{2z}$   
(3)  $\frac{2y}{3xz}$  (4)  $\frac{3z}{2xy}$
- The number of lone pair of electrons present in 2.56 gm of orthorhombic sulphur will be  
(1) 0.02  $\text{N}_0$  (2) 0.16  $\text{N}_0$   
(3) 0.32  $\text{N}_0$  (4) 0.08  $\text{N}_0$
- A 10.0 g of sample of mixture of calcium chloride and sodium chloride is treated with  $\text{Na}_2\text{CO}_3$  to precipitate the calcium as calcium carbonate. This  $\text{CaCO}_3$  is heated to convert all calcium to CaO and final mass of CaO is 1.62 gms. The % by mass of  $\text{CaCl}_2$  in the original mixture is  
(1) 15.2%  
(2) 32.1%  
(3) 21.8%  
(4) 11.07%
- 1000 ml of oxygen at STP were passed through an ozoniser and the resulting volume was 888 ml at STP. This quantity of ozonised oxygen is passed through excess of KI solution. The weight of  $\text{I}_2$  liberated is  
(1) 25.4 g (2) 12.7 g  
(3) 2.54 g (4) 1.27 g
- $x_1$  ml solution having normality 1 is diluted to get a solution of normality  $y_2$ . What would be volume required for dilution?  
(1)  $x_1 \left(1 - \frac{y_1}{y_2}\right)$  (2)  $x_1 \left(\frac{y_2}{y_1} - 1\right)$   
(3)  $x_1 \left(\frac{y_1}{y_2} - 1\right)$  (4)  $x_1 \left(1 + \frac{y_1}{y_2}\right)$
- 1.625 g of an interhalogen compound  $\text{ICl}_x$  was subjected to electric discharge which split it into constituents and brought to STP. It occupied 112 ml. The interhalogen compound is  
(1)  $\text{ICl}$  (2)  $\text{ICl}_3$   
(3)  $\text{ICl}_5$  (4)  $\text{ICl}_7$
- If 12.25 g of a sample of  $\text{KClO}_3$  of 75% purity and rest vapourisable impurities is strongly heated to give  $\text{KCl}$ ,  $\text{KClO}_4$  and 0.32 gm  $\text{O}_2$ . The mass of residue obtained is  
(1) 1.769 g (2) 7.096 g  
(3) 5.327 g (4) 8.85 g
- A 2 g sample of xenon reacts with fluorine. The mass of compound produced is 3.158 g. The empirical formula of the compound is  
(1)  $\text{XeF}_2$  (2)  $\text{XeF}_4$   
(3)  $\text{XeF}_6$  (4)  $\text{XeF}_8$
- 1 g of carbonate of metal was dissolved in 25 ml of 1 N HCl. The resulting liquid required 5 ml of 1 N NaOH for neutralisation. The equivalent weight

# Physical Chemistry Exam Questions

**Dr. Tapas Kumar Das**



## Physical Chemistry Exam Questions:

Physical Chemistry Kenneth S Schmitz, 2018-06-14 The advancements in society are intertwined with the advancements in science To understand how changes in society occurred and will continue to change one has to have a basic understanding of the laws of physics and chemistry Physical Chemistry Multidisciplinary Applications in Society examines how the laws of physics and chemistry physical chemistry explain the dynamic nature of the Universe and events on Earth and how these events affect the evolution of society multidisciplinary applications The ordering of the chapters reflects the natural flow of events in an evolving Universe Philosophy of Science the basis of the view that natural events have natural causes Cosmology the origin of everything from the Big Bang to the current state of the Universe Geoscience the physics and chemistry behind the evolution of the planet Earth from its birth to the present Life Science the molecules and mechanisms of life on Earth Ecology the interdependence of all components within the Ecosphere and the Universe Information Content emphasis on how words and phrases and framing of issues affect opinions reliability of sources and the limitations of knowledge Addresses the four Ws of science Why scientists believe Nature works the way it does Who helped develop the fields of science What theories of natural processes tell us about the nature of Nature and Where our scientific knowledge is taking us into the future Gives a historical review of the evolution of science and the accompanying changes in the philosophy of how science views the nature of the Universe Explores the physics and chemistry of Nature with minimal reliance on mathematics Examines the structure and dynamics of the Universe and our Home Planet Earth Provides a detailed analysis of how humans as members of the Ecosphere have influenced and are continuing to influence the dynamics of events on the paludarium called Earth Presents underlying science of current political issues that shape the future of humankind Emphasizes how words and phrases and framing of issues can influence the opinions of members of society Makes extensive use of metaphors and everyday experiences to illustrate principles in science and social interactions

**Exam Survival Guide: Physical Chemistry** Jochen Vogt, 2017-02-24 A text and exercise book for physical chemistry students This book deals with the fundamental aspects of physical chemistry taught at the undergraduate level in chemistry and the engineering sciences in a compact and practice oriented form Numerous problems and detailed solutions offer the possibility of an in depth reflection of topics like chemical thermodynamics and kinetics atomic structure and spectroscopy Every chapter starts with a recapitulation of important background information before leading over to representative exercises and problems Detailed descriptions systematically present and explain the solutions to the problems so that readers can carefully check their own solutions and get clear cut introductions on how to approach similar problems systematically The book addresses students at the upper undergraduate level as well as tutors and teachers It is a rich source of exercises for exam preparation and can be used alongside classical textbooks Furthermore it can serve teachers and tutors for the conception of their lessons Its well thought through presentation structure and design make the book appeal to everybody who wants to succeed with the

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*AQA AS/A Level Year 1 Chemistry Student Guide: Physical chemistry 1* Alyn G. McFarland, Nora Henry, 2015-10-09 Exam Board AQA Level AS A level Subject Chemistry First Teaching September 2015 First Exam June 2016 Written by experienced examiners Alyn McFarland and Nora Henry this Student Guide for Chemistry Helps you identify what you need to know with a concise summary of the topics examined in the AS and A level specifications Consolidates understanding with tips and knowledge check questions Provides opportunities to improve exam technique with sample answers to exam style questions Develops

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**Understanding Physical Chemistry ,1964 Mosby's Review Questions for the NCLEX-RN Exam - E-Book**

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**Chemical News and Journal of Physical Science** ,1881 **AQA A-level Year 2 Chemistry Student Guide: Inorganic and organic chemistry 2** Alyn G. McFarland,Nora Henry,2016-11-28 Exam Board AQA Level A level Subject Chemistry First Teaching



September 2016 First Exam June 2017 Written by experienced examiners Alyn McFarland and Nora Henry this Student Guide for Chemistry Identifies the key content you need to know with a concise summary of topics examined in the A level specifications Enables you to measure your understanding with exam tips and knowledge check questions with answers at the end of the guide Helps you to improve your exam technique with sample answers to exam style questions Develops your independent learning skills with content you can use for further study and research

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