



Student Exploration: Populations and Samples

Background: Frequency distributions describe the entire population, whereas

Sample Knowledge: Sample distributions describe only a portion of the

population. **Frequency distributions** calculate descriptive statistics such as the mean, standard deviation, and range of values. **Sample distributions** are graphs that represent only a portion of the population. The **frequency distribution** shows the frequency distribution of the entire population, while the **sample distribution** shows the frequency distribution of a portion of the population.

1. How would you describe the shape of this distribution?

Answer: The shape of the distribution is bell-shaped, symmetric, and unimodal. The distribution is centered at the mean.



2. On the above graph, sketch what you think the frequency distribution might look like if the mean height of the population were 100 cm. Label the mean.



Sample Knowledge:

Frequency distributions describe the entire population, while sample distributions describe only a portion of the population. The **frequency distribution** shows the frequency distribution of the entire population, while the **sample distribution** shows the frequency distribution of a portion of the population.



3. How would you describe the shape of the distribution of the frequency distribution of the population? Label the mean.

Answer: The distribution is bell-shaped, symmetric, and unimodal. The distribution is centered at the mean.

4. How would you describe the shape of the distribution of the frequency distribution of the sample? Label the mean.

Answer: The distribution is bell-shaped, symmetric, and unimodal. The distribution is centered at the mean.



Populations And Samples Gizmo Answers

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web in today s experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$ the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a ph regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $\text{Ca(OH)}_2 \rightarrow \text{Ca}^{2+} + 2\text{OH}^{-}$ $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ $2 \times 10^{-2} \text{ M}$ hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^{-}$

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web the chemical equation used was $\text{Ca(OH)}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$ the solubility product was determined to be $K_{\text{sp}} = [\text{Ca}^{2+}][\text{OH}^{-}]^2$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{\text{sp}} = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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experiment a titration of calcium hydroxide with hydrochloric acid

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web question determining the ksp of calcium hydroxide Ca(OH)_2 s Ca^{2+} aq 2OH^- aq for this experiment the titration of calcium hydroxide with hydrochloric acid was investigated determining the ksp of calcium hydroxide Ca(OH)_2 s Ca^{2+} aq 2OH^- aq for this experiment the titration of calcium

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web question chem 182 experiment 8 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water solution of Ca(OH)_2 is represented in equation form solution of calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous as shown below

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for calcium hydroxide was the determination through pH method 2 introduction calcium hydroxide is a strong base soluble in water

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web calculate the ksp for calcium hydroxide Ca(OH)_2 s Ca^{2+} aq 2OH^- aq $K_{sp} = 5 \times 10^{-6}$ find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error accepted value for ksp for sodium hydroxide is 5×10^{-6}

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