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E Durkheim



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web the coefficient in a balanced equation is an idea the concept of telling the chemist that if the atoms molecules and compounds are balanced there are balanced amounts of the atoms molecules and compounds on the other side in the product

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 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $4\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$
 $6\text{AlBr}_3 + 3\text{K}_2\text{SO}_4 \rightarrow 4\text{KBr} + \text{Al}_2\text{SO}_4 + 3\text{Br}_2$
 $4\text{CH}_4 + \text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$
 $8\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

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 $1\text{ N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $4\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$
 $6\text{AlBr}_3 + 3\text{K}_2\text{SO}_4 \rightarrow 4\text{KBr} + \text{Al}_2\text{SO}_4 + 3\text{Br}_2$
 $4\text{CH}_4 + \text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$
 $8\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

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 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $4\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$
 $6\text{AlBr}_3 + 3\text{K}_2\text{SO}_4 \rightarrow 4\text{KBr} + \text{Al}_2\text{SO}_4 + 3\text{Br}_2$
 $4\text{CH}_4 + \text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$

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 coefficients equal to one 1 do not need to be shown in your answers
 $2\text{Fe} + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_3$
 $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
 $2\text{FeBr}_3 + 3\text{H}_2\text{SO}_4 \rightarrow 1\text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2\text{O} + 2\text{C}_4\text{H}_6\text{O}_3$
 $1\text{H}_2\text{O}_2 + 2\text{C}_2\text{H}_4\text{O}_2 \rightarrow$

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 $38\text{Fe} + 2\text{O}_3 \rightarrow 3\text{CO}_2 + 2\text{Fe}_3\text{CO}_2 + 39\text{H}_2 + 3\text{PO}_4 + 3\text{BaO}$
 $3\text{Ba} + 3\text{PO}_4 + 2\text{H}_2\text{O} + 40\text{MgNH}_4\text{PO}_4 + 2\text{Mg} + 2\text{P}_2\text{O}_7 + 2\text{NH}_3 + \text{H}_2\text{O} + 41\text{Al} + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Al}_2\text{SO}_4 + 3\text{H}_2 + 42\text{P}_4\text{O}_{10} + 3\text{HCl} + \text{POCl}_3 + 3\text{HPO}_3 + 43\text{HNO}_2 + 2\text{H}_2\text{O} + 3\text{O}_2 + 4\text{HNO}$

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must satisfy to be consistent with the law of conservation of matter

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web jan 4 2019 a balanced chemical equation gives the number and type of atoms participating in a reaction the reactants products and direction of the reaction balancing an unbalanced equation is mostly a matter of making certain mass and charge are balanced on the reactants and products side of the reaction arrow

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