

Oxidation and Reduction

When happens when elements are transformed in a chemical reaction?

What?

Below reactions show the process by which an element with either properties as the iron, Copper gets oxidized in dissolved gases present in the air. Metals react or corrode in the presence of dissolved water. Sulfuric acids, concentrated HCl and other substances are responsible for the corrosion of metals and also some of the metals are present in the form of compounds. All of these substances are examples of oxidizing and reducing, otherwise known as redox reactions. In this section you will explore when is happening in the above listed in redox reactions.

Model 1 – Redox Reactions

Redox Reactions:



Redox Reactions:



1. What are typical reactions you observe in Model 1?

Redox reactions and corrosion reactions.

2. Explain the redox and corrosion reactions in Model 1. What does the transformation in the redox reactions show about the redox process? Identify them as redox reactions? If yes, what specific changes does Model 1 or supports your answer?

There is no change in the oxidation state of the metal, which is the only change in the redox reactions.

Pogil Lab Oxidation Answer Key

DP Hallahan

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

POGIL Lab Oxidation Answer Key:

Analytical Chemistry Juliette Lantz, Renée Cole, The POGIL Project, 2014-12-31 An essential guide to inquiry approach instrumental analysis Analytical Chemistry offers an essential guide to inquiry approach instrumental analysis collection The book focuses on more in depth coverage and information about an inquiry approach This authoritative guide reviews the basic principles and techniques Topics covered include method of standard the microscopic view of electrochemistry calculating cell potentials the BerriLambert atomic and molecular absorption processes vibrational modes mass spectra interpretation and much more *Oxidation and Reduction* ,1992 **The Mechanisms of Oxidation-reduction**

Reactions in Solution Norman Sutin, 1967 *Oxidation Numbers of Fractionally Charged Atoms* Los Alamos National Laboratory, G. Zweig, K.S. Lackner, 1983 **The Oxidation of Aqueous Potassium Sulfite Solution** Kuan-in Shiau, 1975

The Oxidation of Aldehydes by Fehlings Solution Elizabeth H. Palme, 1971 *Oxidation, in the Gas-phase and in Solution, and with Reactions of Cyclopropane* Charles Frank Howlett Tipper, 1959 Oxidation and Reduction in Inorganic and Analytical Chemistry Alan Vincent, 1985-11-22 Beginning with basic principles this self instructional text leads students to an advanced understanding of oxidation and reduction Covers use of the mole concept and equation balancing and elementary thermodynamics concepts in a clear accessible manner Self tests and revision notes follow each chapter *Keys to Oxidation-reduction* ,1974

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