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POGIL- Stoichiometry

How do chemists use balanced chemical equations?

Why?

Chemists use balanced chemical equations as a basis to calculate how much reactant is needed or product is formed in a reaction. This is called **Stoichiometry**- (stoi-key-ah-meh-tree) Another way of looking at it is using the mole ratio from the balanced equation and information about one compound in the reaction to determine information about another compound in the equation. A mole ratio is a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles. In chemical calculations, mole ratios are used to convert between moles one thing and moles of another.

Purpose: In this activity we will address the question: *How do I convert between different chemical species in a given reaction?*

MODEL 1: Mole-to-Mole Stoichiometry (1-step)

Example: mole-to-mole

Question: Given the following equation:



How many moles of H_2 are needed to produce 34.8 moles of NH_3 ?

Mole ratio



$$\frac{34.8 \text{ mol NH}_3}{1} \times \frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3} = 52.2 \text{ mol H}_2$$

Answer:

Key Questions:

1. What was the given in the example above?
2. Where did $\frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3}$ come from?
3. Double check to see if 52.2 moles of H_2 is the correct answer. SHOW the math below.
4. What units are mole ratios used to convert between? What should you write with the unit mol to distinguish between moles of one substance and another to ensure proper unit cancellation?

*What happened to
Avogadro when he got
bit by 6.02×10^{23}
mosquitoes?*

He got Mol-aria

Pogil Chemistry Mole

Chao Zhang



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Table of Contents Pogil Chemistry Mole

1. Understanding the eBook Pogil Chemistry Mole
 - The Rise of Digital Reading Pogil Chemistry Mole
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Chemistry Mole
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Chemistry Mole
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Chemistry Mole
 - Personalized Recommendations
 - Pogil Chemistry Mole User Reviews and Ratings
 - Pogil Chemistry Mole and Bestseller Lists

5. Accessing Pogil Chemistry Mole Free and Paid eBooks
 - Pogil Chemistry Mole Public Domain eBooks
 - Pogil Chemistry Mole eBook Subscription Services
 - Pogil Chemistry Mole Budget-Friendly Options
6. Navigating Pogil Chemistry Mole eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Chemistry Mole Compatibility with Devices
 - Pogil Chemistry Mole Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Chemistry Mole
 - Highlighting and Note-Taking Pogil Chemistry Mole
 - Interactive Elements Pogil Chemistry Mole
8. Staying Engaged with Pogil Chemistry Mole
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Chemistry Mole
9. Balancing eBooks and Physical Books Pogil Chemistry Mole
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Chemistry Mole
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pogil Chemistry Mole
 - Setting Reading Goals Pogil Chemistry Mole
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pogil Chemistry Mole
 - Fact-Checking eBook Content of Pogil Chemistry Mole
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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