

An Introduction to the Ideal Gas Law

$$PV = nRT$$

$$P_1 V_1 = n_1 R_1 T_1$$

$$P_2 V_2 = n_2 R_2 T_2$$

$$\frac{P_1 V_1}{n_1 T_1} = R_1 = R_2 = \frac{P_2 V_2}{n_2 T_2}$$

$$\frac{P_1 V_1}{\cancel{n_1 T_1}} = \frac{P_2 V_2}{\cancel{n_2 T_2}}$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Combined
Gas Law

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

Charles'
Law

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$

Gay-Lussac's
Law

$$\frac{V_1}{n_1} = \frac{V_2}{n_2}$$

Avogadro's
Law

$$P_1 V_1 = P_2 V_2$$

Boyle's Law

If you know which variables are held constant
ANY gas law can be derived!!!

Pogil Chemistry Gas Laws

Marco Cascella



Pogil Chemistry Gas Laws:

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