

Bond Energy

What makes a reaction endothermic or exothermic?

Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** (ΔH). In this activity we will look at one way energy changes can be approximated for chemical reactions.

Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol _{rxn})		Reaction	Change in Energy (kJ/mol _{rxn})
A)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
 - When bonds are broken (~~formed~~) there is a positive energy change.
 - Breaking bonds is endothermic (~~exothermic~~).
 - When bonds are broken (~~formed~~) there is a negative energy change.
 - Forming bonds is exothermic (~~endothermic~~).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.
 - How do the changes in energy for the reverse reactions compare?
same absolute value, opposite sign
 - Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.

amt of energy required to break is same as amount released when formed

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O García

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Pogil Bond Energy Answer Key

1. Understanding the eBook Pogil Bond Energy Answer Key
 - The Rise of Digital Reading Pogil Bond Energy Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Bond Energy Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Bond Energy Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Bond Energy Answer Key
 - Personalized Recommendations
 - Pogil Bond Energy Answer Key User Reviews and Ratings
 - Pogil Bond Energy Answer Key and Bestseller Lists
5. Accessing Pogil Bond Energy Answer Key Free and Paid eBooks
 - Pogil Bond Energy Answer Key Public Domain eBooks
 - Pogil Bond Energy Answer Key eBook Subscription Services
 - Pogil Bond Energy Answer Key Budget-Friendly Options
6. Navigating Pogil Bond Energy Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Bond Energy Answer Key Compatibility with Devices
 - Pogil Bond Energy Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Bond Energy Answer Key
 - Highlighting and Note-Taking Pogil Bond Energy Answer Key
 - Interactive Elements Pogil Bond Energy Answer Key
8. Staying Engaged with Pogil Bond Energy Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Pogil Bond Energy Answer Key
- 9. Balancing eBooks and Physical Books Pogil Bond Energy Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Bond Energy Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Bond Energy Answer Key
 - Setting Reading Goals Pogil Bond Energy Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Bond Energy Answer Key
 - Fact-Checking eBook Content of Pogil Bond Energy Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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