

## Free Energy

Where does the energy come from that keeps all living things alive?

### Why?

All living things require a constant flow of energy. This is one of the big ideas in science. If the flow of energy to an ecosystem is cut-off, like an extended lack of sunlight or food, then all the living things in the ecosystem will die. But you have heard of the Law of Conservation of Energy, so why wouldn't there always be enough energy to go around?

### Model 1 – Spontaneous Processes

|   | Process Description                                                           | Change in Enthalpy ( $\Delta H$ ) | Change in Entropy ( $\Delta S$ ) | Spontaneous?                        |
|---|-------------------------------------------------------------------------------|-----------------------------------|----------------------------------|-------------------------------------|
| A | Two pure substances $\rightarrow$ Homogeneous mixture                         | -                                 | Increasing                       | Yes                                 |
| B | Solute on one side of membrane $\rightarrow$ Solute on both sides of membrane | -                                 | Increasing                       | Yes                                 |
| C | Polypeptide chain $\rightarrow$ Individual amino acids                        | Exothermic (negative)             | Increasing (positive)            | Yes                                 |
| D | $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$                                     | Exothermic                        | Increasing                       | Yes                                 |
| E | $6CO_2 + 6H_2O \rightarrow C_6H_{12}O_6 + 6O_2$                               | Endothermic                       | Decreasing                       | No                                  |
| F | Glucose $\rightarrow$ Starch                                                  | Endothermic                       | Decreasing                       | No                                  |
| G | Water $\rightarrow$ Ice                                                       | Exothermic                        | Decreasing                       | Below 0 °C, yes<br>Above 0 °C, no   |
| H | Cold water (25 °C) $\rightarrow$ Hot water (60 °C)                            | Endothermic                       | Increasing                       | Below 60 °C, no<br>Above 60 °C, yes |

1. Consider Model 1.

a. What two terms are used to describe the enthalpy changes for the reactions?

b. What is the symbol for the change in enthalpy?

c. What sign (+ or -) does the change in enthalpy have when the reaction releases heat energy to the surroundings?

# Pogil Biology Answer Key Energy

**Arshad Iqbal**



## **Pogil Biology Answer Key Energy:**

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