

PHYSICAL: Kinetic Molecular Theory

Learning Objectives

- At standard atmospheric conditions, all gases have approximately the same kinetic energy, but are made of different molecules and are made up of different numbers of particles. (The kinetic molecular theory is based on the assumption that all particles have the same average kinetic energy.)
- At standard atmospheric conditions, all gases have approximately the same average kinetic energy.

Model 1: Representation of Gases in Different Phases

The following figure illustrates the representation of particles in three phases: solid, liquid, and gas.

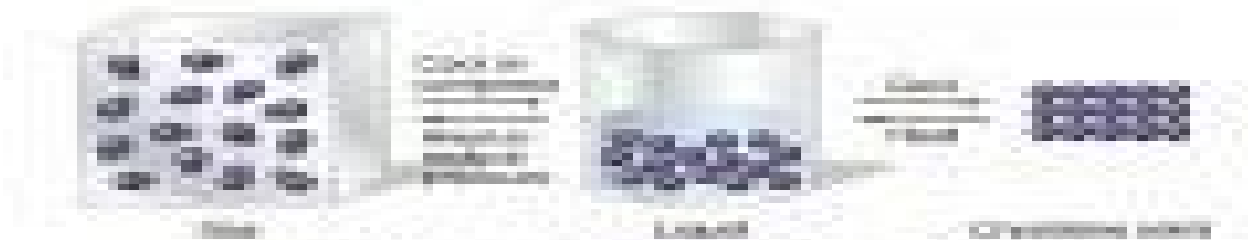


Figure 1: Representation of particles in different phases of matter

Model 1: Representation of particles in different phases of matter. The particles are represented by small circles. The particles in the solid phase are closely packed and arranged in a regular pattern. The particles in the liquid phase are closely packed and arranged in a disordered pattern. The particles in the gaseous phase are widely spaced and arranged in a disordered pattern.

Key Concepts

1. All gases have the same average kinetic energy at standard atmospheric conditions. This is because the kinetic energy of a particle is proportional to the square of its velocity, and the average velocity of all particles in a gas is the same.

Table 1: Representation of particles in different phases

	Solid	Liquid	Gaseous
How are the particles arranged?	The particles are arranged in a regular pattern. They are closely packed together. The intermolecular forces are strong.	The particles are arranged in a disordered pattern. They are closely packed together. The intermolecular forces are moderate.	The particles are arranged in a disordered pattern. They are widely spaced. The intermolecular forces are weak.
How do the particles move?	The particles vibrate in fixed positions. They have low kinetic energy.	The particles move in a disordered manner. They have moderate kinetic energy.	The particles move in a disordered manner. They have high kinetic energy.

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