

Reaction Quotient

How do you predict which direction a reaction will proceed to reach equilibrium?

Why?

When a reaction reaches equilibrium there must be some non-negligible amount of every species in the reaction, otherwise the reaction cannot react in both directions. Knowing this, it is very easy to predict which direction a reaction will go to reach equilibrium when one of the components of the reaction has an initial concentration of zero. Many of the problems you have worked with thus far have some component at zero concentration, but real life does not work that way. Most of the time, the reaction in question has some measurable amount of every species. Deciding which way the reaction will go to reach equilibrium then becomes more challenging.

Model 1 – A Theoretical Equilibrium

Trial 1	$A(g) + B(g) \rightleftharpoons C(g)$		
Initial	1.000 M	1.000 M	1.000 M
Change			
Equilibrium	1.464 M	1.464 M	0.536 M

Trial 3	$A(g) + B(g) \rightleftharpoons C(g)$		
Initial	1.000 M	0.500 M	1.500 M
Change			
Equilibrium	1.864 M	1.364 M	0.636 M

Trial 5	$A(g) + B(g) \rightleftharpoons C(g)$		
Initial	1.400 M	1.200 M	0.400 M
Change			
Equilibrium	1.388 M	1.188 M	0.412 M

Trial 2	$A(g) + B(g) \rightleftharpoons C(g)$		
Initial	2.000 M	0.500 M	0.500 M
Change			
Equilibrium	2.150 M	0.650 M	0.350 M

Trial 4	$A(g) + B(g) \rightleftharpoons C(g)$		
Initial	1.600 M	1.000 M	0.400 M
Change			
Equilibrium	1.600 M	1.000 M	0.400 M

Trial 6	$A(g) + B(g) \rightleftharpoons C(g)$		
Initial	0.750 M	2.000 M	0.250 M
Change			
Equilibrium	0.675 M	1.925 M	0.325 M

- Examine Model 1.
 - Write the theoretical chemical reaction that is used in the trials of Model 1.
 - If 0.50 M of reactant A reacts, predict the change in concentration of B and C.
- What variables were changed in the different trials shown in Model 1?

Pogil Activities For Ap Chemistry Answer Key

Thomas Griffiths



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Table of Contents Pogil Activities For Ap Chemistry Answer Key

1. Understanding the eBook Pogil Activities For Ap Chemistry Answer Key
 - The Rise of Digital Reading Pogil Activities For Ap Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Activities For Ap Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Activities For Ap Chemistry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Activities For Ap Chemistry Answer Key

- Personalized Recommendations
 - Pogil Activities For Ap Chemistry Answer Key User Reviews and Ratings
 - Pogil Activities For Ap Chemistry Answer Key and Bestseller Lists
5. Accessing Pogil Activities For Ap Chemistry Answer Key Free and Paid eBooks
 - Pogil Activities For Ap Chemistry Answer Key Public Domain eBooks
 - Pogil Activities For Ap Chemistry Answer Key eBook Subscription Services
 - Pogil Activities For Ap Chemistry Answer Key Budget-Friendly Options
 6. Navigating Pogil Activities For Ap Chemistry Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Activities For Ap Chemistry Answer Key Compatibility with Devices
 - Pogil Activities For Ap Chemistry Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Activities For Ap Chemistry Answer Key
 - Highlighting and Note-Taking Pogil Activities For Ap Chemistry Answer Key
 - Interactive Elements Pogil Activities For Ap Chemistry Answer Key
 8. Staying Engaged with Pogil Activities For Ap Chemistry Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Activities For Ap Chemistry Answer Key
 9. Balancing eBooks and Physical Books Pogil Activities For Ap Chemistry Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Activities For Ap Chemistry Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Pogil Activities For Ap Chemistry Answer Key
 - Setting Reading Goals Pogil Activities For Ap Chemistry Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Pogil Activities For Ap Chemistry Answer Key

- Fact-Checking eBook Content of Pogil Activities For Ap Chemistry Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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