

Calculating pH

How will it be used for the concentration differences here?

Why?

As buffers and other substances control pH in a system, it is a way to regulate the acidity or basicity of a cell. Also, this property is also controlled through pH levels in sports as athletes regulate their acid/base balance as indicated pH values that relate to various things, including the amount of pH in a system, before it could lead down to a specific value. When the concentration difference (pH) is greater than the amount.

Table 1 – Ion Concentrations for Acids and Bases

Acids	Substance	Acidic, Basic or Neutral?	$[H^+]$ (M)	$[OH^-]$ (M)	$[H^+] > [OH^-]$?
1	0.1 M HCl (strong)	Acidic	$1.0 \times 10^{-1} M$	$1.0 \times 10^{-13} M$	Yes
2	0.001 M HCl (strong)	Acidic	$1.0 \times 10^{-3} M$	$1.0 \times 10^{-11} M$	Yes
3	0.00001 M HCl (strong)	Acidic	$1.0 \times 10^{-5} M$	$1.0 \times 10^{-9} M$	Yes
4	0.000001 M HCl (strong)	Acidic	$1.0 \times 10^{-6} M$	$1.0 \times 10^{-8} M$	Yes
5	0.0000001 M HCl (strong)	Neutral	$1.0 \times 10^{-7} M$	$1.0 \times 10^{-7} M$	Yes
6	0.00000001 M HCl (strong)	Basic	$1.0 \times 10^{-8} M$	$1.0 \times 10^{-6} M$	Yes
7	0.000000001 M HCl (strong)	Basic	$1.0 \times 10^{-9} M$	$1.0 \times 10^{-5} M$	Yes
8	0.0000000001 M HCl (strong)	Basic	$1.0 \times 10^{-10} M$	$1.0 \times 10^{-4} M$	Yes
9	0.00000000001 M HCl (strong)	Basic	$1.0 \times 10^{-11} M$	$1.0 \times 10^{-3} M$	Yes
10	0.000000000001 M HCl (strong)	Basic	$1.0 \times 10^{-12} M$	$1.0 \times 10^{-2} M$	Yes

1. What are the results of pH? (Acidic, Basic, Neutral?)

Acidic, Basic, Neutral?

2. For Buffer 1, which one has a higher concentration, hydrochloric or hydrobromic?

Hydrochloric

3. How did you determine the concentration of hydrochloric and how did you determine the Buffer 1 and Buffer 2 acids, the concentration of the acid?

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Pogil Activities Calculating Ph Answers

Robert R. Redfield



Pogil Activities Calculating Ph Answers:

Decoding **Pogil Activities Calculating Ph Answers**: Revealing the Captivating Potential of Verbal Expression

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