

Balancing Chemical Equations

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STANDARD 1 LEARNING MODULE 10

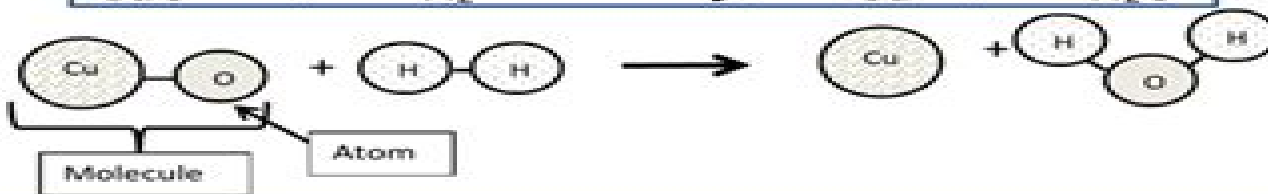
Background Information and Why: Chemical reactions can be described by chemical equations. Equations means equal. One each side of the equal side, also known as "yields", the number AND type of atoms are exactly the same. The Law of Conservation of Mass states that the atoms cannot be created or destroyed. The atoms can only be rearranged into different molecules/substances. Atoms are the building blocks of molecules. Therefore, in order to obey The Law of Conservation of Mass, the number of atoms for each element must be the same in the reactants (left side) as it is in the products (right side). In order to balance equations we cannot split or add to existing molecules. We can change the coefficient, adding molecules. We cannot change the subscripts, the small numbers, because they determine how many atoms are in a molecule.

REACTION A

Chemical Equation



Atomic Representation



REACTION B

Chemical Equation



Atomic Representation



Questions

0.) Label one more molecule and one more atom in EACH reaction. Use the ones done above to help you. Pick another atom draw an arrow to it. Label it "atom". Then circle or "bracket" another molecule and label it "molecule".

1.) How many of each type of atom are represented in each side of reaction A? Count them up and fill out the table below.

# Atoms in Reactants	Atoms in Products
Cu- ONE	Cu- ONE
O- ONE	O- ONE
H- TWO	H- TWO

Pogil Types Of Chemical Reactions Balancing Equations

Michael Rajnik



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