

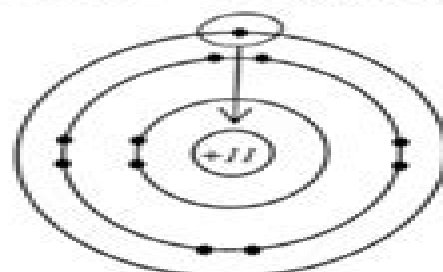
Advanced Periodic Trends

How do the particles in an atom interact to dictate periodic properties of elements?

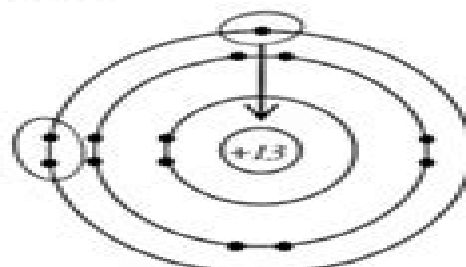
Why?

Previously you have learned about Coulombic attraction and how it governs the trends in properties among the elements of the Periodic Table. For example, as you move across a period, atoms have more protons in the nucleus, which pulls the electrons in tighter making smaller, more electronegative atoms. As you proceed down a column of the periodic table, the valence electrons get further away, and the atoms get larger and less electronegative. Inquisitive students will always ask, "If more protons in the nucleus pull electrons in tighter, then why don't atoms get smaller when going down a column?" That is a good question, and one that you will investigate in this activity.

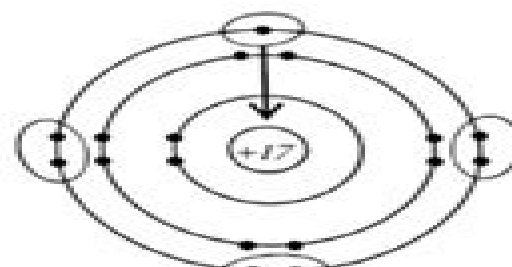
Model 1 – Period 3 Elements



Sodium



Aluminum



Chlorine

Atomic Number	11	13	17
Number of Electrons	11	13	17
Number of Valence Electrons	1	3	7
Core Charge	+1	+3	+7
Atomic Radius	186 pm	143 pm	99 pm
1st Ionization Energy	496 kJ/mol	578 kJ/mol	1251 kJ/mol
Electronegativity	0.9	1.5	3.0

1. Describe the relationship of the three elements in Model 1 with regard to their relative positions on the periodic table.

The three elements are in the same row of the periodic table.

2. Refer to a periodic table to complete the first three rows of the table in Model 1.

See Model 1.

Pogil Periodic Trends

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Table of Contents Pogil Periodic Trends

1. Understanding the eBook Pogil Periodic Trends
 - The Rise of Digital Reading Pogil Periodic Trends
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Periodic Trends
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Periodic Trends
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Periodic Trends
 - Personalized Recommendations
 - Pogil Periodic Trends User Reviews and Ratings
 - Pogil Periodic Trends and Bestseller Lists
5. Accessing Pogil Periodic Trends Free and Paid eBooks

- Pogil Periodic Trends Public Domain eBooks
- Pogil Periodic Trends eBook Subscription Services
- Pogil Periodic Trends Budget-Friendly Options
- 6. Navigating Pogil Periodic Trends eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Periodic Trends Compatibility with Devices
 - Pogil Periodic Trends Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Periodic Trends
 - Highlighting and Note-Taking Pogil Periodic Trends
 - Interactive Elements Pogil Periodic Trends
- 8. Staying Engaged with Pogil Periodic Trends
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Periodic Trends
- 9. Balancing eBooks and Physical Books Pogil Periodic Trends
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Periodic Trends
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Periodic Trends
 - Setting Reading Goals Pogil Periodic Trends
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Periodic Trends
 - Fact-Checking eBook Content of Pogil Periodic Trends
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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