

POGIL: Lewis Dot Structures Part 1 (20 P)

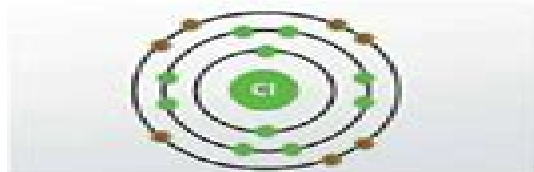
How are electrons shared to create covalently bonded molecules?

Read This!

Why? Covalent bonds result from a *sharing* of electrons by two or more atoms (usually nonmetals). Lewis theory (Gilbert Newton Lewis, 1875-1946) focuses on the valence electrons, since the outermost electrons are the ones that are highest in energy and farthest from the nucleus, and are therefore the ones that are most exposed to other atoms when bonds form. Lewis dot diagrams for elements are a handy way of picturing valence electrons, and especially, which electrons are available to be shared in covalent bonds.

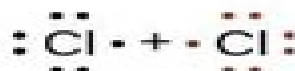
Bohr Model: Chlorine

7 valence electrons on outermost energy level



Lewis Dot diagram: Chlorine

7 valence electrons



chlorine
atoms



chlorine
molecule



Lewis Dot Diagrams Showing Valence Electrons for Individual Elements

1	2		3	4	5	6	7	8
		Transition metals						

Atoms are grouped together by number of valence electrons; all atoms in Group 1 have 1 valence electron.

Pogil On Lewis Structures

JL Elias



Pogil On Lewis Structures:

Unveiling the Magic of Words: A Review of "**Pogil On Lewis Structures**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "**Pogil On Lewis Structures**," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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