Naming Molecular Compounds

Why?

How are the chemical formula and name of a molecular compound related?

When you began chemistry class this year, you probably already knew that the chemical formula for car-bon dioxide was CO. Today you will find out why CO, is named that way. Naming chemical compounds correctly is of paramount importance. The slight difference between the names carbon monoxide (CO, a

poisonous, deadly gas) and carbon dioxide (CO, a greenhouse gas that we exhale when we breathe out) can be the difference between life and death! In this activity you will learn the naming system for molecular compounds.

Model 1 - Molecular Compounds

Moleculur Formula	Number of Atoms of First	Number of Atoms of Second	Name of Compound	
CLF	1	1	Chlorine monofhonde	
CIF,	1	5	Chlorine pentafhronide	
CO	1	1	Carbon monoxide	
CO ₂	1	5	Carbon dioxide	
C1,0	2	1	Dichlorine monoxide	
PC1,	1	5	Phosphoras pertachlroride	
N,O,	2	5	Dinitrogen pentoxide	

- Fill in the table to indicate the number of atoms of each type in the molecular formula.
- Examine the molecular formulas given in Model 1 for various molecular compounds.
 - a. How many different elements are present in each compound shown?

Two

b. Do the compounds combine metals with metals, metals with nonmetals, or nonmetals with nonmetals?

Non-metals with non-metals

c. Based on your answer to b, what type of bonding must be involved in molecular compounds?

Covalent bonds

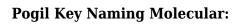
Find all of the compounds in Model 1 that have chlorine and fluorine in them. Explain why
the name "chlorine fluoride" is not sufficient to identify a specific compound.

The number of fluorine atoms is different in the two compounds, but name does not reflect the difference

4. Assuming that the name of the compound gives a clue to its molecular formula, predict how many atoms each of these prefixes indicates, and provide two examples.

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Table of Contents Pogil Key Naming Molecular

- 1. Understanding the eBook Pogil Key Naming Molecular
 - The Rise of Digital Reading Pogil Key Naming Molecular
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Pogil Key Naming Molecular
 - Exploring Different Genres
 - $\circ\,$ Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Key Naming Molecular
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Pogil Key Naming Molecular
 - Personalized Recommendations

- Pogil Key Naming Molecular User Reviews and Ratings
- Pogil Key Naming Molecular and Bestseller Lists
- 5. Accessing Pogil Key Naming Molecular Free and Paid eBooks
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 - Pogil Key Naming Molecular eBook Subscription Services
 - Pogil Key Naming Molecular Budget-Friendly Options
- 6. Navigating Pogil Key Naming Molecular eBook Formats
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 - Pogil Key Naming Molecular Compatibility with Devices
 - Pogil Key Naming Molecular Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Pogil Key Naming Molecular
 - Highlighting and Note-Taking Pogil Key Naming Molecular
 - o Interactive Elements Pogil Key Naming Molecular
- 8. Staying Engaged with Pogil Key Naming Molecular
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Key Naming Molecular
- 9. Balancing eBooks and Physical Books Pogil Key Naming Molecular
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Pogil Key Naming Molecular
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Key Naming Molecular
 - Setting Reading Goals Pogil Key Naming Molecular
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Key Naming Molecular
 - Fact-Checking eBook Content of Pogil Key Naming Molecular

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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