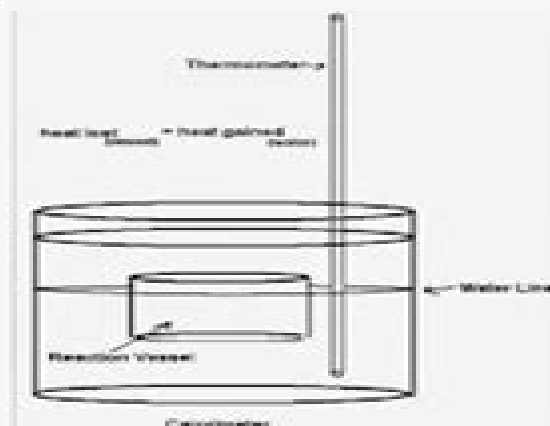


Model

Assume that a calorimeter is a closed system where all the energy released by an exothermic change is absorbed by the water in the calorimeter. If the mass of the water is known, the temperature change of the water can be used to determine the amount of heat energy released.



Equation for the calculation of heat is: $q = mC\Delta T$

q = heat released in Joules

m = mass of water in the calorimeter

C = specific heat capacity (**not $^{\circ}\text{C}$ which is a temperature**)

ΔT = [final temperature – initial temperature] (absolute value)

Key Questions:

1. What is the numerical value and units of the specific heat capacity of water?

Specific heat for water is $4.18 \frac{\text{J}}{\text{g}^{\circ}\text{C}}$ or $1 \frac{\text{cal}}{\text{g}^{\circ}\text{C}}$

2. What information does the specific heat capacity of water provide?

The specific heat value gives the amount of heat necessary to raise the temperature of one gram of water up 1°C .

3. How can the heat released into some mass of water be calculated from the specific heat capacity of water and the change in temperature of the water? Answer in words not with an equation.

Specific heat is how how much heat is needed to raise water 1°C , so if you how much the temperature changed, you can multiply the ΔT by specific heat to get the heat need to create temperature change.

Pogil Heat Temperature And Calorimetry Answers

Jicheng Xie



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