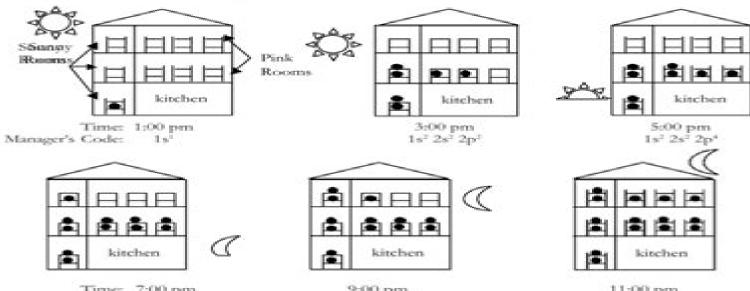
Electron Configurations

What is the electron structure in an atom?

Why?

The electron structure of an atom is very important. Scientists use the electronic structure of atoms to predict bonding in molecules, the charge(s) an atom might have, and the physical properties of elements. In order for scientists to describe the electron structure in an atom, they give the electrons "addresses." Just like your address might include your house number, street, city, and state, an electron's "address" has multiple parts. In this activity, you will learn how the electrons fill up the available spaces in an atom and how their "addresses" or configurations are assigned.

Model 1 - The Boarding House



Time: 7:00 pm Manager's Code: 1s² 2s² 2ps 3s¹ 9:00 pm 1s² 2s² 2p⁶ 3s² 3p⁴

- 11:00 pm 1s² 2s² 2p⁴ 3s² 3p³
- Examine the boarding house diagrams in Model 1. Match each symbol below with its most likely meaning.
 - _^ +

- I. Bunk bed for boarders
- c. 1s2 2s2 2p4 3s1
- II. Manager's code for the number of boarders in the house and their room assignments.
- III. Boarder

Pogil Activities For High School Chemistry Electron Configurations

Brendan G. Carr

Pogil Activities For High School Chemistry Electron Configurations:

Structures and Approximations for Electrons in Molecules David B. Cook,1978 Electron Structure and High-temperature Chemistry of Coordination Compounds, 1996 *Introduction to the Electron Theory of Small Molecules* Atomic and Molecular Electronic Configuration Revisited Harold J. Teague, 2014-06-11 In Andrew Crowther Hurley, 1976 AMEC chapter one you will find a structure for O2 supported by both bond order and unpaired electron data Previous attempts at providing an electronic structure for O2 fails due to bond order number of unpaired electrons or oxygen atoms indicating five orbitals although four is the limit Herein you will find the newly described LONE P ORBITAL Pi BOND which is the solution not only for O2 but also for NO and OF uncharacterized The LPOP bond system is possible due to the small size and high electronegative character of the NO and F atoms which makes it related to H Bonding Diatomic structures involving row two periodic table elements LiBe Be2 BeB B2 BC C2 CN N2 and F2 are also presented All structures are shown being formed by two methods MRAE Most Recently Added Electron and LCAO Linear Combination of Atomic Orbitals Chapter one also includes a Section III MRAE Mechanistic Sequence For NO NO O2 OF F2 and F2 F F0 Section IV Photographs of NO and O2 molecular models Section V MRAE MO Energy Level diagrams for the diatomic structures listed above and Section VI a MRAE MO Energy Level Diagrams Template which is based on the cascade of changes in electronic distribution when an electron is forced onto a previous structure MRAE such as N2 going to N2 and NO going to NO Chapter 2 includes sections on hybridization and mechanism as it relates to polyatomic species. The following sequences are given C H CH H CH2 H CH3 H CH4 saturated and tetrahedral arrangement and N H NH H NH2 H NH3 H NH4 also saturated and tetrahedral arrangement The emphasis is to show trend from atomic structure through sp sp2 and sp3 hybridization Chapter 3 is composed of selections from Atomic and Molecular Configuration 1987 Included is MRAE AUFBAU TEMPLATE of the periodic table elements expanded to two pages It is informative to compare the MRAE concept for atoms to the MRAE method for diatomic species in Chapter one In both cases the key is the changes in electronic structure when an electron is Chemistry Univ 1 Archie Gibb, Arthur A. Sandison, Learning and Teaching Scotland, Andrew Watson, 2000 added

Chemistry David John Cole-Hamilton, Iain Patterson, 2001 Electron Configuration 148 Success Secrets - 148

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Poincare Ion Formation of polyatomic and molecular ions Valence electrons Copernicium Physical and atomic Cadmium Mol Phys Term symbol Valence electron Electron configuration Pnictogen Chemical Electronic configuration Electron configuration in molecules Group 11 elements Hund s rules Octet rule Explanation in quantum theory Group 11 element Characteristics Hassium Physical and atomic Alkali metal Extended periodic table Superactinides Neutron absorption Neutron absorbers Flerovium Nuclear stability and isotopes Lewis acid Comparison with Bronsted Lowry Theory Ununoctium Calculated atomic and physical properties Tunnel diode F block Isotopes of strontium Ytterbium Aufbau principle Atoms Energy levels Pauli exclusion principle Atoms and the Pauli principle Electron shell List of elements with electrons per shell Periodic table Layout Rutherfordium Chemical Antimony Group 8 element Lanthanum Chemical properties Extended periodic table Eka superactinides Manganese Promethium Physical properties and much more

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