

Name \_\_\_\_\_

# Physical Science Chemical Bonding Study Guide

## A. Covalent & Ionic Compounds/Bonds

1. Which ion always forms with a charge of 2-? \_\_\_\_\_ (anion)  
which cation is always a 2+? \_\_\_\_\_ (cation)
2. Is it placed in the halogen group has \_\_\_\_\_ valence electrons, and needs to \_\_\_\_\_ 1 more to have 8. This is because it has \_\_\_\_\_ valence.
3. Is it \_\_\_\_\_ bond to the adjacent elements oppositely charged ions.  
PCl<sub>3</sub> is covalent bond is \_\_\_\_\_
4. Is it considered ionic, atomic, \_\_\_\_\_, molecular, formed by an ionic bond, atomic, \_\_\_\_\_, molecular.
5. 2-bonded group? Which side bond has shared completely by it completely bond is to \_\_\_\_\_, and others group's shared equally to the \_\_\_\_\_  
How can predict if is covalent, ionic, polar covalent, or ionic bond calculated by electronegativity? \_\_\_\_\_
6. Is it atomic or ionic bond has \_\_\_\_\_ valence electrons. This family contains an ionic particular bond that all have 8 valence electrons is called \_\_\_\_\_
7. Atomic ions gain, lose, or share \_\_\_\_\_ by becoming more stable.

## B. Formulas for Ionic Compounds – Put in the right number

| Write correct coefficient | Magnesium | Fluorine | Phosphorus | Chlorine |
|---------------------------|-----------|----------|------------|----------|
| Electrons                 |           |          |            |          |
| Covalent bond             |           |          |            |          |

3. What is the chemical formula for each of the following ions:  
compounds?

Show your work either above the Lewis structures, and use arrows to show the transfer of electrons. OR use the periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if they calculate numbers already given.

# Physical Science Chemical Bonding Study Guide

**AN Whitehead**



## **Physical Science Chemical Bonding Study Guide:**

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the changes within the main text and the labeling and coloring exercises are taken from the illustrations designed for the book Filled with empowering self study tools and learning activities for every learning style this practical Study Guide follows the organization of the main text chapter by chapter helping students every step of the way toward content mastery The variety of learning activities with three main components are designed to facilitate student learning of all aspects of anatomy physiology and the effects of disease not merely to test knowledge

### **TEExES Core Subjects EC-6 (391) Book + Online**

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James E. Brady,Fred Senese,2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH<sub>4</sub> from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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Appropriations. Subcommittee on Departments of Labor, and Health, Education, and Welfare, and Related Agencies, 1965

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