

# Light & Periodicity HOMEWORK

## Lesson 1: Homework

1. Select the correct term, related to EACH symbol below:

- c.  $2.998 \times 10^8$   
 d. wavelength  
 e. frequency  
 f. energy  
 g.  $2.998 \times 10^{17}$

- a.  $6.626 \times 10^{-34}$   
 b. frequency  
 c. energy  
 d.  $2.998 \times 10^8$   
 e. Wavelength

2. Identify the units for EACH symbol below:

- a.  $3.00 \times 10^8$   
 b.  $\text{m}^{-1}$   
 c.  $\text{Joule}^{-1}$

- a. meters (m)  
 b. 1/photon  
 c.  $\text{sec}^{-1}$   
 d. nanometers (nm)  
 e. 1/m  
 f. 1/second

3. Convert  $6.65 \times 10^2$  nanometers to meters.

$$6.65 \times 10^2 \text{ nm} \times \frac{1 \text{ m}}{10^9 \text{ nm}} = 6.65 \times 10^{-7} \text{ m}$$

4. At some point during their orbit around the sun, Mars is  $2.48 \times 10^8$  meters from Earth.

a) If a signal, which uses light waves, is sent from Mars to Earth (a wavelength of  $2.00 \times 10^2$  meters), what is its frequency?

$$\text{wavelength} \\ (2.00 \times 10^2) \times \text{frequency} = (2.48 \times 10^8) \\ \rightarrow \text{frequency} = \frac{2.48 \times 10^8}{2.00 \times 10^2} = 1.24 \times 10^6 \text{ sec}^{-1}$$

b) The units on the frequency is  $\text{sec}^{-1}$ . If this value is inverted, the units are inverted as well. How many seconds does it take to send this signal? Give your answer with units from part "a" in this calculation.

$$\frac{1}{1.24 \times 10^6 \text{ sec}^{-1}} = 8.06 \times 10^{-7} \text{ sec}$$

5. Light can be used to excite the electrons in atoms. Certain frequencies of light work best for specific atoms. To excite an atom of Calcium, the ideal frequency of  $6.2 \times 10^{14} \text{ s}^{-1}$  is needed. What is the wavelength (in nm) of this electromagnetic radiation? Give your answer in scientific notation and with units.

$$\text{wavelength} \\ (6.2 \times 10^{14}) \times \text{wavelength} = (3.00 \times 10^8) \\ \rightarrow \text{wavelength} = \frac{3.00 \times 10^8}{6.2 \times 10^{14}} = 4.84 \times 10^{-7} \text{ m}$$

$$4.84 \times 10^{-7} \text{ m} \times \frac{10^9 \text{ nm}}{1 \text{ m}} = 484 \text{ nm}$$

6. The energy required to dissociate the  $\text{Cl}_2$  molecule into two separate Cl atoms requires a wavelength of  $5.00 \times 10^2$  meters. What is the energy of this light in  $\text{Joules}^{-1}$ ?

$$\text{wavelength} \\ (5.00 \times 10^2) \times \text{frequency} = (3.00 \times 10^8) \\ \rightarrow \text{frequency} = \frac{3.00 \times 10^8}{5.00 \times 10^2} = 6.00 \times 10^5 \text{ sec}^{-1}$$

$$\text{energy} \\ (6.00 \times 10^5 \text{ sec}^{-1}) \times (6.626 \times 10^{-34} \text{ Joules} \cdot \text{sec}) = 3.9756 \times 10^{-28} \text{ Joules}$$

7. Refrigerator coolant contains special compounds known as chlorofluorocarbons (CFCs). If not properly disposed, a refrigerator can leak this gas into the environment. Thus, ultraviolet radiation from the sun can change these CFCs, which can then destroy the ozone layer. How much energy do the CFCs receive from UV light that has a frequency of  $6.60 \times 10^{14} \text{ s}^{-1}$ . Give your answer in  $\text{Joules}^{-1}$ .

$$\text{energy} \\ (6.60 \times 10^{14} \text{ sec}^{-1}) \times (6.626 \times 10^{-34} \text{ Joules} \cdot \text{sec}) = 4.37316 \times 10^{-19} \text{ Joules}$$

8. You want your lawn sprinklers to come on in the morning, at first light. So, you design a switch that will operate by the photoelectric effect when enough sunlight hits it. As sunlight falls on your switch, the metal inside the device absorbs energy and ejects electrons to create electricity that powers the sprinklers. If the switch requires  $5.0 \times 10^{-19} \text{ Joules}$  of energy to remove an electron, will the switch work if the sunlight has a wavelength of 540nm?

$$\text{wavelength} \\ (540 \text{ nm}) \times \text{frequency} = (540 \times 10^{-9} \text{ m}) \times \text{frequency} = 5.4 \times 10^{-7} \text{ m} \times \text{frequency}$$

$$\text{energy} \\ (5.4 \times 10^{-7} \text{ m}) \times \text{frequency} = (5.4 \times 10^{-7} \text{ m}) \times \text{frequency} = 5.4 \times 10^{-7} \text{ m} \times \text{frequency}$$

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**CH Cherryholmes**



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