

## Periodic Trends Worksheet

*Directions: Use your notes to answer the following questions.*

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- |  |                  |
|--|------------------|
| a. atomic radius (excluding noble gases) | <b>Decreases</b> |
| b. first ionization energy               | <b>Increases</b> |
| c. electronegativity                     | <b>Increases</b> |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **decreases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

# Periodic Trends Lab Answer Key

**Pam Walker, Elaine Wood**



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*The Good Student* Charles Jay Rop, 1996 *The Best of ICCAD* Andreas Kuehlmann, 2012-12-06 In 2002 the International Conference on Computer Aided Design ICCAD celebrates its 20th anniversary This book commemorates contributions made by ICCAD to the broad field of design automation during that time The foundation of ICCAD in 1982 coincided with the growth of Large Scale Integration The sharply increased functionality of board level circuits led to a major demand for more powerful Electronic Design Automation EDA tools At the same time LSI grew quickly and advanced circuit integration

became widely available. This in turn required new tools using sophisticated modeling analysis and optimization algorithms in order to manage the evermore complex design processes. Not surprisingly during the same period a number of start up companies began to commercialize EDA solutions complementing various existing in house efforts. The overall increased interest in Design Automation DA required a new forum for the emerging community of EDA professionals one which would be focused on the publication of high quality research results and provide a structure for the exchange of ideas on a broad scale. Many of the original ICCAD volunteers were also members of CANDE Computer Aided Network Design a workshop of the IEEE Circuits and Systems Society. In fact it was at a CANDE workshop that Bill McCalla suggested the creation of a conference for the EDA professional. Bill later developed the name *The World of Chemistry* Jones, Jones Staff, 1996.

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