

Oxidation-Reduction Reactions

- Electrons can be neither created out of nothing nor destroyed
- In any redox reaction there is an element being reduced and an element being oxidized
- The total increase in the oxidation numbers must equal the total decrease in the oxidation numbers

If an element is reduced

- It gains electrons
- Oxidation number decreases
- The substance is classified as an oxidizing agent

If an element is oxidized

- It loses electrons
- Oxidation number increases
- The substance is classified as a reducing agent

Oxidation And Reduction Key

R Bogdan

Oxidation And Reduction Key:

Regents Exams and Answers: Chemistry--Physical Setting Revised Edition Barron's Educational Series, Albert Tarendash, 2021-01-05 Barron's Regents Exams and Answers Chemistry provides essential practice for students taking the Chemistry Regents including actual recently administered exams and thorough answer explanations for all questions This book features Eight actual administered Regents Chemistry exams so students can get familiar with the test Thorough explanations for all answers Self analysis charts to help identify strengths and weaknesses Test taking techniques and strategies A detailed outline of all major topics tested on this exam A glossary of important terms to know for test day

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