

POGIL 2 Ions

How are ions made from neutral atoms?

You have learned that not all atoms of an element are the same. Variation in the number of neutrons results in different isotopes of the element. In this activity we will explore another variation that can take place—the loss and gain of electrons. The exchange of electrons between atoms is a very common way for chemical change to take place. We will see it many times throughout the year.

1. Use Diagram 1 on the next page to complete the following table.

	Metal or Nonmetal	Number of protons in the atom	Number of protons in the ion	Number of neutrons in the atom	Number of neutrons in the ion	Number of electrons in the atom	Number of electrons in the ion	Charge on the ion
Lithium	metal	3	3	4	4	3	2	1+
Magnesium	<i>Metal</i>	12	12	12	12	12	10	2+
Aluminum	<i>Metalloid</i>	13	13	14	14	13	10	3+
Fluorine	<i>Non metal</i>	9	9	10	10	9	10	1-
Oxygen	nonmetal	8	8	8	8	8	10	2-
Nitrogen	<i>Non metal</i>	7	7	7	7	7	10	3-

2. Based on the table you completed in Question 1, what distinguishes a neutral atom from an ion?

A neutral atom has the same number of protons and electrons (no charge). An ion has a charge (# of protons does not equal # of electrons).

3. Examine the isotope symbols in Diagram 1.

- a. Where is the ion charge located in the isotope symbol?

Top right corner

- b. Is a charge indicated on the neutral atoms? If yes, where is it located?

No

4. Which subatomic particle carries a positive charge?

Protons are positive

5. Which subatomic particle carries a negative charge?

Electrons are negative

6. Propose a mathematical equation to calculate the charge on an ion from the number of protons and electrons in an ion. Confirm that your equation works using two positive ion examples and two negative ion examples from Diagram 1.

The charge of an ion = # of protons - # of electrons

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Table of Contents Pogil Packet Answer Keys

1. Understanding the eBook Pogil Packet Answer Keys
 - The Rise of Digital Reading Pogil Packet Answer Keys
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Packet Answer Keys
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Packet Answer Keys
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Packet Answer Keys
 - Personalized Recommendations
 - Pogil Packet Answer Keys User Reviews and Ratings
 - Pogil Packet Answer Keys and Bestseller Lists
5. Accessing Pogil Packet Answer Keys Free and Paid eBooks

- Pogil Packet Answer Keys Public Domain eBooks
- Pogil Packet Answer Keys eBook Subscription Services
- Pogil Packet Answer Keys Budget-Friendly Options
- 6. Navigating Pogil Packet Answer Keys eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Packet Answer Keys Compatibility with Devices
 - Pogil Packet Answer Keys Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Packet Answer Keys
 - Highlighting and Note-Taking Pogil Packet Answer Keys
 - Interactive Elements Pogil Packet Answer Keys
- 8. Staying Engaged with Pogil Packet Answer Keys
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Packet Answer Keys
- 9. Balancing eBooks and Physical Books Pogil Packet Answer Keys
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Packet Answer Keys
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Packet Answer Keys
 - Setting Reading Goals Pogil Packet Answer Keys
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Packet Answer Keys
 - Fact-Checking eBook Content of Pogil Packet Answer Keys
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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