

Names _____
Period _____

Heat, Temperature & Calorimetry

Part 1: Heat & Temperature

Heat deals with energy flow. **Temperature** is a ratio of energy per molecule. The two are very different. For instance, a Fourth-of-July sparkler emits sparks with temperatures hovering around 2000 °C, but the actual amount of heat received when a spark lands on ones hand or face is very small. High temperature means a high ratio of energy per molecule.

Heat is measured in units of energy, most commonly the *joule*. Heat is also commonly measured in calories or Calories.

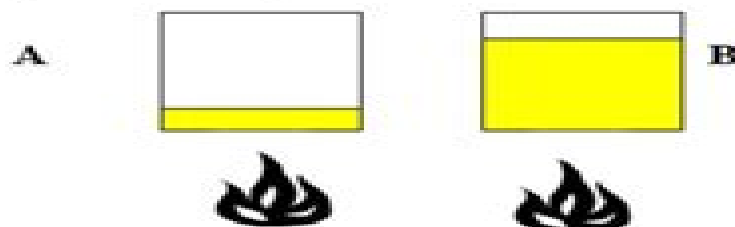
Temperature is measured in degrees. Temperature is most commonly measured in degrees *Celsius*.

Introduction:

You have already been introduced to the formal definitions of heat and temperature. The purpose of this exercise is to explore these concepts as they relate to some physical applications.

Part 1:

You place identical pots on the stove and heat them equally. Pot A contains 1.00 g of water;



pot B contains 100g of water. Both pots start out at 20°C.

1. If you wanted to heat both pots of water until they boiled, which pot would boil first? Why?
2. When both pots are boiling what is the temperature in A? B?
3. Compare the amount of heat contained by the water in each pot. Are they the same or different? If they are different explain why and indicate which one you may think has more (or less).

Pogil Heat Temperature Answer Key

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