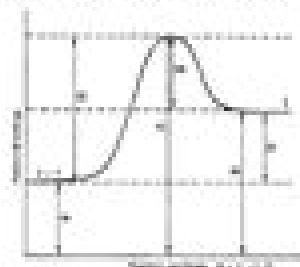
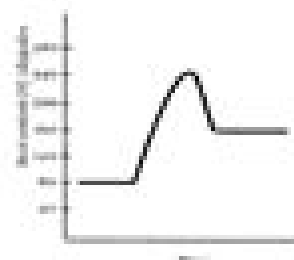


Potential Energy Diagram Worksheet



- Which of the letters a-f in the diagram represents the potential energy of the products? e
- Which letter indicates the potential energy of the activated complex? c
- Which letter indicates the potential energy of the reactants? a
- Which letter indicates the activation energy? b
- Which letter indicates the heat of reaction? f
- Is the reaction endothermic or exothermic? Endo
- Which letter indicates the activation energy of the reverse reaction? d
- Which letter indicates the heat of reaction of the reverse reaction? f
- Is the reverse reaction endothermic or exothermic? Exo



- The heat content of the reactants of the forward reaction is about 80 kilojoules.
- The heat content of the products of the forward reaction is about 20 kilojoules.
- The heat content of the activated complex of the forward reaction is about 160 kilojoules.
- The activation energy of the forward reaction is about 80 kilojoules.
- The heat of reaction (ΔH) of the forward reaction is about -60 kilojoules.
- The forward reaction is Exo (endothermic or exothermic).
- The heat content of the reactants of the reverse reaction is about 20 kilojoules.
- The heat content of the products of the reverse reaction is about 80 kilojoules.
- The heat content of the activated complex of the reverse reaction is about 160 kilojoules.
- The activation energy of the reverse reaction is about 140 kilojoules.
- The heat of reaction (ΔH) of the reverse reaction is about +60 kilojoules.
- The reverse reaction is Exo (endothermic or exothermic).

Key Energy 915 #1 - Reaction Rates

- Chemical reactions occur when reactants collide. For what reasons may a collision fail to produce a chemical reaction?

Not enough energy or wrong angle

- If every collision between reactants lead to a reaction, what determines the rate at which the reaction occurs?

of effective collisions

- What is the activation energy of a reaction, and how is this energy related to the activated complex of the reaction?

Ea is the energy needed to form the A.C.

- What happens when a catalyst is used in a reaction?

Catalysts lower the Ea, thus speeding up reactions

- Name 4 things that will speed up or slow down a chemical reaction.

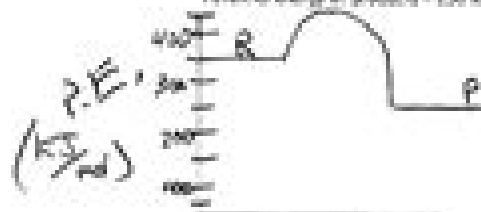
- ΔT Temp
- ΔP Press.
- C
- Catalyst

- Draw an energy diagram for a reaction. (label the axis)

Potential energy of reactants = 350 kJ/mole

Activation energy = 200 kJ/mole

Potential energy of products = 250 kJ/mole



Time

- Is the reaction in #6 exothermic or endothermic? Explain.

Exo, because P has less E than R

- How could you lower the activation energy for the reaction in #6?

Use a catalyst.

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D Siedentop



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