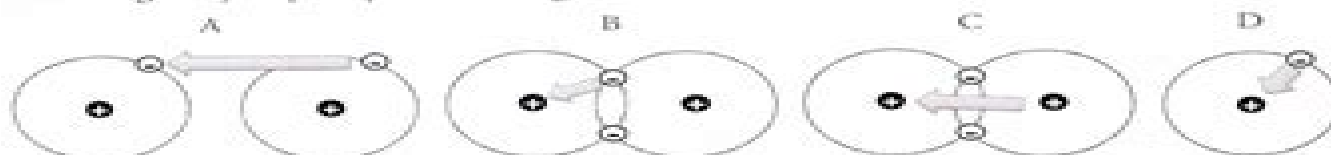


Read This!

Electronegativity is a measure of the ability of an atom's nucleus to attract electrons from a different atom within a covalent bond. A higher electronegativity value correlates to a stronger pull on the electrons in a bond. This value is only theoretical. It cannot be directly measured in the lab.

12. Using the definition stated in the *Read This!* box above, select the best visual representation for electronegativity. Explain your reasoning.



13. Locate the electronegativity values in Model 1.

- a. What is the trend in electronegativity going down a group in Model 1?

Electronegativity decreases as you go down a group (vertical column).

- b. Explain the existence of the trend described in part a in terms of atomic structure and Coulombic attraction.

The ability to attract electrons is based upon attractive force. The greater the distance between the nucleus and valence electrons, the less the attractive force. Less force means a lower electronegativity.

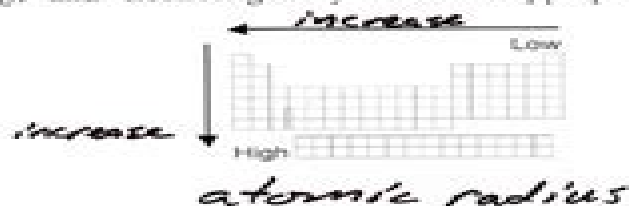
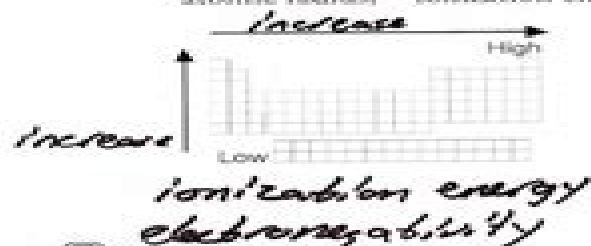
- c. What is the trend in electronegativity going across a period in Model 1?

Electronegativity increases from left to right across a period (horizontal row)

- d. Explain the existence of the trend described in part c in terms of atomic structure and Coulombic attraction.

The greater the number protons in an atom, in a particular period, the greater the attractive force between the nucleus and valence electrons. The greater attractive force means a higher electronegativity.

14. The two diagrams below can summarize each of the three trends discussed in this activity. Write "atomic radius," "ionization energy," and "electronegativity" under the appropriate diagram.



Pogil Answer Key Chemistry Periodic Trends

R Sanford



Pogil Answer Key Chemistry Periodic Trends:

CBSE Class 11 Chemistry Chapter 5 Story in English Harishkumar Soni,Priti Singhal,2024-09-30 Chapter 5 of CBSE

Class 11 Chemistry explores the historical development and significance of the periodic table The chapter begins with early efforts to categorize elements leading to Mendeleev s groundbreaking periodic table which arranged elements by atomic mass It highlights the modern periodic table organized by atomic number and introduces key concepts such as periods groups and periodic trends in properties like ionization energy electronegativity and atomic radius This understanding is vital for predicting the behavior of elements in chemical reactions and emphasizes the systematic nature of chemistry **Atoms**

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