

5. Examine the chemical formulas for the Arrhenius bases in Model 1. What anion do all the Arrhenius base chemical formulas have in common?

contain OH!

6. A student dissolved a small amount of baking soda in water and tested it with litmus paper. The litmus paper turned blue. Is baking soda likely an acid or a base?

BASE!

Read This!

In 1903 Svante Arrhenius won the Nobel Prize in Chemistry for defining acids and bases in terms of the ions produced. An Arrhenius acid is any substance that produces hydrogen ions [or hydronium ions (H_3O^+) a hydrogen ion attached to a water molecule] when dissolved in water. An Arrhenius base is any substance that produces hydroxide ions when dissolved in water. While the Arrhenius definitions of acids and bases is useful, it is limited. Johannes Brønsted and Thomas Lowry developed more general definitions for acids and bases using H^+ ion (proton) transfer as the focus.

Model 2 – Brønsted-Lowry Acids and Bases

Reaction 1	HCl(g) (acid)	+	$\text{H}_2\text{O(l)}$ (base)	$\rightleftharpoons$	$\text{H}_3\text{O}^+(\text{aq})$ + $\text{Cl}^-(\text{aq})$	gives accepts
Reaction 2	$\text{NH}_3(\text{aq})$ (base)	+	HF(aq) (acid)	$\rightleftharpoons$	$\text{NH}_4^+(\text{aq})$ + $\text{F}^-(\text{aq})$	gives accepts
Reaction 3	$\text{NH}_4^+(\text{aq})$ (acid)	+	$\text{H}_2\text{O(l)}$ (base)	$\rightleftharpoons$	$\text{NH}_3(\text{aq})$ + $\text{H}_3\text{O}^+(\text{aq})$	gives accepts
Reaction 4	$\text{F}^-(\text{aq})$ (base)	+	$\text{H}_2\text{PO}_4^-(\text{aq})$ (acid)	$\rightleftharpoons$	$\text{HPO}_4^{2-}(\text{aq})$ + HF(aq)	gives accepts

7. Identify the Brønsted-Lowry acids in Model 2.

a. Atoms of which element are present in all of the Brønsted-Lowry acids in Model 2?

H

b. How can you tell from Reaction 1 that HCl loses an H^+ ion rather than a hydrogen atom when the reaction occurs? *Hint: Look at the products.*

8. For each acid-base reaction in Model 2, describe the role of the Brønsted-Lowry acid in the H^+ ion (proton) transfer that occurs.

9. For each acid-base reaction in Model 2, describe the role of the Brønsted-Lowry base in the proton (H^+ ion) transfer that occurs.

Pogil Acids And Bases Answer Key

Siddappa N.Byraredddy



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