

Intermolecular Forces Notes

Section 2

What is the Lewis Dot Structure for Carbon? Based on this Lewis Dot Structure, how many electrons can Carbon bond with?



Carbon only has two valence electrons that are represented available for bonding. So, carbon should only be able to bond with TWO atoms.

So how do we explain the chemical CH_4 ? Two theories: (1) Valence Bond Theory (2) Molecular Orbital Theory

Valence Bond Theory (VBT Theory)

It is it carbon can hybridize, meaning they mix together.

Orbital Diagram of Carbon:



Here, all 2s orbitals and 2 p orbitals have hybridized. The orbitals sp hybridization.

Types of hybridization: (1) sp (2) sp² (3) sp³

"A "connection" can be a BOND (line) or a LONE PAIR (2 dots)

Practice Problem:

What type of hybridization occurs in the molecule PF_3 ?



Since the molecule has THREE connections around P, it's sp³ hybridization.



Methane: CH_4

(2) Methane: CH_4

Practice Problem:

What type of hybridization, if any, occurs in the following molecules?

(1) BF_3

(2) BCl_3



sp² around B



sp² around B

Valence Bond Theory includes a trick

A single bond is called a sigma bond. Its symbol is σ.

A double bond is 2 sigma & 2 pi bonds. Its symbol is π.

A triple bond contains 1 sigma and 2 pi bonds.



Practice Problem:

How many sigma bonds are in this molecule of acetone? How many pi bonds?

3 sigma
1 pi



Practice Problem:

Use valence bond theory to describe the bonding in N_2 .



Each nitrogen has sp hybridization.
There is 1 sigma & 2 pi bonds.

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