Balancing redox reactions under Basic Conditions $Cr(OH)_3 + ClO_3^{-1} --> CrO_4^{2-} + Cl^{-1} \text{ (basic)}$

- Lets balance the reduction one first
- for every Oxygen add a water on the other side
- For every hydrogen add a H⁺ to the other side
- Each H* will react with an OH on both sides (whatever you do to one side...you do the same to the other)
- H⁺ and OH⁻ make water
- cancel the waters
- Balance the imbalance of charge with electrons (-1 vs. -7)

- ClO₃ --> Cl-
- ClO₃⁻ --> Cl⁻ + 3H₂O
- 6H⁺ + ClO₃⁻ --> Cl⁻ + 3H₂O
- 6 OH⁻ + 6H⁺ + ClO₃⁻ --> Cl⁻ + 3H₂O + 6
 OH⁻
- 6H₂O + ClO₃ --> Cl + 3H₂O + 6 OH
- 3H₂O + ClO₃⁻ --> Cl⁻ + 6 OH⁻
- 6e⁻ + 3H₂O + ClO₃⁻ --> Cl⁻ + 6 OH⁻

Pogil Balancing Redox Equations

E Staff

Pogil Balancing Redox Equations:

Whispering the Techniques of Language: An Emotional Quest through Pogil Balancing Redox Equations

In a digitally-driven world where screens reign great and quick communication drowns out the subtleties of language, the profound strategies and emotional nuances concealed within words often go unheard. However, set within the pages of **Pogil Balancing Redox Equations** a fascinating fictional treasure pulsing with organic emotions, lies a fantastic quest waiting to be undertaken. Penned by an experienced wordsmith, this charming opus invites viewers on an introspective trip, softly unraveling the veiled truths and profound affect resonating within the cloth of each and every word. Within the mental depths of this poignant review, we will embark upon a sincere exploration of the book is key styles, dissect their charming publishing fashion, and succumb to the effective resonance it evokes serious within the recesses of readers hearts.

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