

POGIL 2 Ions

How are ions made from neutral atoms?

You have learned that not all atoms of an element are the same. Variation in the number of neutrons results in different isotopes of the element. In this activity we will explore another variation that can take place—the loss and gain of electrons. The exchange of electrons between atoms is a very common way for chemical change to take place. We will see it many times throughout the year.

1. Use Diagram 1 on the next page to complete the following table.

	Metal or Nonmetal	Number of protons in the atom	Number of protons in the ion	Number of neutrons in the atom	Number of neutrons in the ion	Number of electrons in the atom	Number of electrons in the ion	Charge on the ion
Lithium	metal	3	3	4	4	3	2	1+
Magnesium	<i>Metal</i>	12	12	12	12	12	10	2+
Aluminum	<i>Metalloid</i>	13	13	14	14	13	10	3+
Fluorine	<i>Non metal</i>	9	9	10	10	9	10	1-
Oxygen	nonmetal	8	8	8	8	8	10	2-
Nitrogen	<i>Non metal</i>	7	7	7	7	7	10	3-

2. Based on the table you completed in Question 1, what distinguishes a neutral atom from an ion?

A neutral atom has the same number of protons and electrons (no charge). An ion has a charge (# of protons does not equal # of electrons).

3. Examine the isotope symbols in Diagram 1.

- a. Where is the ion charge located in the isotope symbol?

Top right corner

- b. Is a charge indicated on the neutral atoms? If yes, where is it located?

No

4. Which subatomic particle carries a positive charge?

Protons are positive

5. Which subatomic particle carries a negative charge?

Electrons are negative

6. Propose a mathematical equation to calculate the charge on an ion from the number of protons and electrons in an ion. Confirm that your equation works using two positive ion examples and two negative ion examples from Diagram 1.

The charge of an ion = # of protons - # of electrons

Pogil Activity Ion Answer Key

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Project**



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Analytical Chemistry Juliette Lantz, Renée Cole, The POGIL Project, 2014-12-31 An essential guide to inquiry approach instrumental analysis Analytical Chemistry offers an essential guide to inquiry approach instrumental analysis collection The book focuses on more in depth coverage and information about an inquiry approach This authoritative guide reviews the basic principles and techniques Topics covered include method of standard the microscopic view of electrochemistry calculating cell potentials the Beer-Lambert atomic and molecular absorption processes vibrational modes mass spectra interpretation and much more

Unveiling the Magic of Words: A Report on "**Pogil Activity Ion Answer Key**"

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Table of Contents Pogil Activity Ion Answer Key

1. Understanding the eBook Pogil Activity Ion Answer Key
 - The Rise of Digital Reading Pogil Activity Ion Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Activity Ion Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Activity Ion Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Activity Ion Answer Key
 - Personalized Recommendations
 - Pogil Activity Ion Answer Key User Reviews and Ratings
 - Pogil Activity Ion Answer Key and Bestseller Lists

5. Accessing Pogil Activity Ion Answer Key Free and Paid eBooks
 - Pogil Activity Ion Answer Key Public Domain eBooks
 - Pogil Activity Ion Answer Key eBook Subscription Services
 - Pogil Activity Ion Answer Key Budget-Friendly Options
6. Navigating Pogil Activity Ion Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Activity Ion Answer Key Compatibility with Devices
 - Pogil Activity Ion Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Activity Ion Answer Key
 - Highlighting and Note-Taking Pogil Activity Ion Answer Key
 - Interactive Elements Pogil Activity Ion Answer Key
8. Staying Engaged with Pogil Activity Ion Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Activity Ion Answer Key
9. Balancing eBooks and Physical Books Pogil Activity Ion Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Activity Ion Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pogil Activity Ion Answer Key
 - Setting Reading Goals Pogil Activity Ion Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pogil Activity Ion Answer Key
 - Fact-Checking eBook Content of Pogil Activity Ion Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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