

Periodic Trends Worksheet

Directions: Use the trends in electronegativity to answer the questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, O, S, Cl

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

K, Na, S, Cl, O

3. Which atom has a higher electronegativity: fluorine or sulfur?

Fluorine is more electronegative than sulfur.

4. Which two elements in the same family generally have similar properties?

There are many elements that share a similar property such as potassium.

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (increases to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest?

Period 7, francium, cesium, barium, potassium, rubidium, sodium, lithium

7. Which period in the periodic table shows atoms with the smallest atomic radius? Which element has the smallest?

Period 1, hydrogen, helium, lithium, beryllium, boron, carbon, nitrogen, oxygen, fluorine, neon

8. Write the atoms in each group that have the largest atomic radius.

a. **Li**, **Na**, **K**
 b. **Be**, **Mg**, **Ca**
 c. **B**, **Al**, **Ga**
 d. **C**, **Si**, **Ge**
 e. **N**, **P**, **As**

f. **O**, **S**, **Se**
 g. **F**, **Cl**, **Br**
 h. **Ne**, **Ar**, **Kr**

9. Circle the atom in each group that has the greatest electronegativity.

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DJ Losen



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