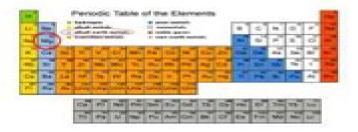
POGIL: Períodic Table Trends

(Adapted from Rush Henrietta CSD)

Why::

The Periodic Table is one of the greatest inventions in the history of man. It allows scientists to predict physical and chemical properties of the elements. Dimitri Mendeleev (a Russian scientist) and Robert Mosley (a British chemist) put together this table in the late 1800's based on properties of the elements known at the time. The trends in these properties as you go across periods and down groups is the subject of this Chem POGIL.



Success Criteria:

- Understand the meaning of
 - atomic radius, reactivity, electronegativity, ionization energy
- Recognize trends in atomic radius, reactivity, ionization energy and electronegativity as you go across
 periods and down groups.

Resources:

Reference Table S, Periodic Table, Unit 3 (Chapter 6) Notes, & Chapter 6 in textbook

Prerequisites:

Answer the following questions using the resources above.

- What happens to the number of valence electrons as you go down a group on the periodic table?
- 2. What happens to the number of valence electrons as you go from left to right across a period?
- 3. What happens to the number of energy levels as you go down a group on the periodic table?
- 4. What happens to the number of energy levels as you go from left to right across a period?
- 5. Draw a Lewis structure for a Cl atom.

- Draw a Bohr diagram for the Cl atom. (Draw above)
- In this lab we will be talking about atomic radius.
 - a. Look up the atomic radius of CI on Table S: _____
 - b. Why is it possible to show the atomic radius using a Bohr diagram but not by using a Lewis structure?
 - On your Bohr diagram draw an arrow to represent the radius of the Cl atom.
- Define periodic:
- 9. Define trend:

Periodic Trends Pogil Answers

Patrick Vollmar

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Table of Contents Periodic Trends Pogil Answers

- 1. Understanding the eBook Periodic Trends Pogil Answers
 - The Rise of Digital Reading Periodic Trends Pogil Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Periodic Trends Pogil Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Pogil Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodic Trends Pogil Answers
 - Personalized Recommendations
 - Periodic Trends Pogil Answers User Reviews and Ratings
 - Periodic Trends Pogil Answers and Bestseller Lists
- 5. Accessing Periodic Trends Pogil Answers Free and Paid eBooks
 - Periodic Trends Pogil Answers Public Domain eBooks
 - Periodic Trends Pogil Answers eBook Subscription Services
 - Periodic Trends Pogil Answers Budget-Friendly Options

- 6. Navigating Periodic Trends Pogil Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Periodic Trends Pogil Answers Compatibility with Devices
 - Periodic Trends Pogil Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Pogil Answers
 - Highlighting and Note-Taking Periodic Trends Pogil Answers
 - Interactive Elements Periodic Trends Pogil Answers
- 8. Staying Engaged with Periodic Trends Pogil Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Pogil Answers
- 9. Balancing eBooks and Physical Books Periodic Trends Pogil Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Pogil Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Pogil Answers
 - Setting Reading Goals Periodic Trends Pogil Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Pogil Answers
 - Fact-Checking eBook Content of Periodic Trends Pogil Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

• Interactive and Gamified eBooks

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